

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode $T \gg \theta_M$) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute $\frac{3}{2}R$ to the molar heat capacity. If we denote the number of active rotational modes by ν_R^\ddagger (so for most molecules at normal temperatures $\nu_R^\ddagger = 2$ for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

- 16.10 The equipartition theorem states that the average energy of a molecule is $\frac{1}{2}kT$ per degree of freedom. For a diatomic molecule, there are 5 degrees of freedom (3 translational and 2 rotational) at room temperature. The average energy per molecule is $\frac{5}{2}kT$. The average energy per mole is $\frac{5}{2}RT$. The average energy per mole is $\frac{5}{2}RT$.

Solutions to exercises

- 16.11 The internal energy of a gas is given by $U = \frac{3}{2}nRT$ for a monatomic gas. The average energy per molecule is $\frac{3}{2}kT$.

$$U = \frac{3}{2}nRT$$

16.12 The internal energy of a gas is given by $U = \frac{3}{2}nRT$ for a monatomic gas.

$$U = \frac{3}{2}nRT$$

- 16.13 The internal energy of a gas is given by $U = \frac{3}{2}nRT$ for a monatomic gas.

$$U = \frac{3}{2}nRT$$

- 16.14 The internal energy of a gas is given by $U = \frac{3}{2}nRT$ for a monatomic gas.

16.15 The internal energy of a gas is given by $U = \frac{3}{2}nRT$ for a monatomic gas.

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