

16.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|---|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. hydronium ion (H ₃ O ⁺) | 6. neutral solution |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|---|-----------------------------------|
| 1. $[H^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 16.1) | 2. $pH = -\log[H^+]$ (Eq. 16.2) |
| 3. $pH = -\log[H^+]$ (Eq. 16.3) | 4. $pOH = -\log[OH^-]$ (Eq. 16.4) |

Part A: Completion

Use the appropriate number of significant figures in the answers to the following questions that are based on the following data: $[H^+] = 1.0 \times 10^{-3} M$ and $[OH^-] = 1.0 \times 10^{-11} M$.

1. What substance is _____ in water? Hydroxide ion (OH⁻)
2. What substance is _____ in the autoionization of water which produces an acidic solution? _____
3. What substance is _____ in water? _____
4. _____ is _____ in water. _____
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19 2 Hydrogen Ions And Acidity Answer Key

Eileen Ramsden



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