

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|--------------------------------|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. pH |
| 4. hydronium ion | 5. hydronium ion concentration | 6. hydronium ion concentration |
| | 7. hydroxide ion | 8. [OH ⁻] |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Write the number in the space provided.

1. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (water autoionization)
2. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces an acidic solution)
3. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a basic solution)
4. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
5. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
6. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
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9. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

10. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

11. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

12. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

13. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

14. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

15. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

16. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

17. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

18. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

19. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

20. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

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H Kauffman



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