

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. hydroxide ion |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 14.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 14.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 14.3) | 4. $pH = -\log[H_3O^+]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction given in the following examples.

1. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (water autoionization)
2. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces an acidic solution)
3. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a basic solution)
4. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
5. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
6. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
7. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
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9. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)
10. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (the autoionization of water which produces a neutral solution)

11. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

12. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

13. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

14. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

15. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

16. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

17. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

18. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

19. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

20. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

21. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

22. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$

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