

Mass of the  $\text{BaSO}_4$  formed =  $x$  gram

233 g of  $\text{BaSO}_4$  contains 32 g of Sulphur.

Therefore,  $x$  gram of  $\text{BaSO}_4$  will contain,

$$= \left( \frac{32}{233} \times \frac{x}{w} \right) \text{g of S}$$

Percentage of Sulphur,

$$= \left( \frac{32}{233} \times \frac{x}{w} \times 100 \right) \%$$

**Solved Example:**

In an estimation of sulphur by carius method, 0.2175g of the substance gave 0.5825g of  $\text{BaSO}_4$  calculate the percentage composition of S in the compound.

**Solution:**

Weight of organic compound 0.2175g

Weight of  $\text{BaSO}_4$  0.5825g

233g of  $\text{BaSO}_4$  contains 32g of S

0.5825g of  $\text{BaSO}_4$  contains  $\frac{32}{233} \times \frac{0.5825}{0.2175}$

Percentage of S =  $\frac{32}{233} \times \frac{0.5825}{0.2175} \times 100$

# Determination Of Sulphur Containing Volume 1

**S Baum**



## **Determination Of Sulphur Containing Volume 1:**

## Reviewing **Determination Of Sulphur Containing Volume 1**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Determination Of Sulphur Containing Volume 1**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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