

**Department of Chemical Engineering**  
**Series Test 1- CH 010 502 Chemical Reaction Engineering**

**Time: 1 hr**

**Maximum Marks:30**

**Part 1**

- 1) At 500 K the rate of a bimolecular reaction is ten times the rate at 400 K. Find the activation energy of this reaction using Collision theory.(2 marks)
- 2) For an exothermic reversible reaction which one of the following correctly describes the dependence of equilibrium constant (K) with temperature (T) and pressure (P) ?
  - a) K is independent of T and P
  - b) K increases with increase in T and P
  - c) K increases with T and decreases with P
  - d) K decreases with an increase in T and is independent of P (1 marks)
- 3) For which reaction order, the half-life of the reactant is half of the full lifetime (time for 100% conversion) of the reactant? (A) Zero order (B) Half order (C) First order (D) Second order (1 marks)
- 4) For the chemical reaction  $P \rightarrow Q$ , it is found that the rate of reaction doubles as the concentration of 'P' is doubled. If the reaction rate is proportional to  $C_p^n$ , then what is the value of 'n' for this chemical reaction ? (A) 1 (B) 2 (C) 3 (D) 0 (1 marks)

**Part 2**

( 3×5 = 15 marks)

- 1) The concentrations of a compound undergoing chemical change were 5.72, 3.23 and 1.96 at the times 0, 20 and 50 min from the commencement of reaction .Ascertain the order of the reaction.
- 2) Discuss various intermediates used in non-elementary chemical reactions
- 3) Briefly explain temperature dependence from Arrhenius theory

**Part 3**

(10 marks)

Answer any one of the following

- 1) The following data are obtained at 0°C in a constant volume batch reactor using pure gaseous A:

Time in min	0	2	4	6	8	10	12	14	$\infty$
pA in mm Hg	760	60	475	390	320	275	240	215	150

The reaction is  $A \rightarrow P$  with  $E_a = 2.5$  K .Find the rate equation for this reaction.

**OR**

- 2) a) Explain Psuedo-steady state hypothesis (PSSH)(3 marks)

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