

Problem 5.2

Calculate the molar volume and the compressibility factor for ethylene at 79.7 °C and 20.2 bar by using the following equations;

- a) The ideal gas equation.
- b) The truncated virial equation with the following experimental values of virial coefficients: $B = -140 \text{ cm}^3/\text{mol}$ and $C = 7500 \text{ cm}^6/\text{mol}^2$
- c) The generalized correlation for the compressibility factor

Solution

(a) $Z = 1$
 $V = RT/P$

$$V = \frac{RT}{P} = \frac{(8.314 \frac{\text{J}}{\text{mol}\cdot\text{K}})(352.85 \text{ K})}{20.2 \times 10^5 \text{ Pa}} = 1.465 \times 10^{-3} \frac{\text{m}^3}{\text{mol}}$$

(b) $Z = PV/RT = 1 + B/V + C/V^2$
 with $P = 20.2 \cdot 10^5 \text{ Pa}$, $R = 8.3145 \text{ J}/(\text{mol}\cdot\text{K})$, $T = 352.85 \text{ K}$,

$$Z = 1 + \frac{B}{V} + \frac{C}{V^2} = 1 + \frac{-140 \times 10^{-6} \text{ m}^3/\text{mol}}{1.465 \times 10^{-3} \text{ m}^3/\text{mol}} + \frac{7500 \times 10^{-12} \text{ m}^6/\text{mol}^2}{(1.465 \times 10^{-3} \text{ m}^3/\text{mol})^2} = 1.054$$

or

$$Z = 1 + \frac{B}{V} + \frac{C}{V^2} = 1 + \frac{-140 \times 10^{-6} \text{ m}^3/\text{mol}}{1.465 \times 10^{-3} \text{ m}^3/\text{mol}} + \frac{7500 \times 10^{-12} \text{ m}^6/\text{mol}^2}{(1.465 \times 10^{-3} \text{ m}^3/\text{mol})^2} = 1.054$$

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Chemical Engineering Rapid Problem Index

JE Gale



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