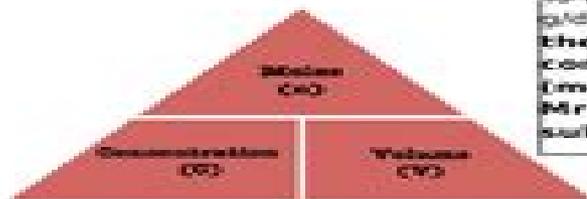


Chemical calculations



Units
 Mass = g
 moles = mol
 Relative formula mass = g mol⁻¹

Moles and concentration

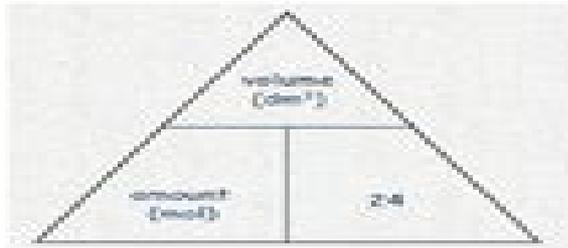


To convert into g/dm³ multiply the concentration (mol/dm³) by the Mr of the substance

Moles = Mass ÷ Relative formula mass
 Mass = Moles × Relative formula mass
 Relative formula mass = Mass ÷ Moles

$$1 \text{ dm}^3 = 1000 \text{ cm}^3$$

Moles and gas volumes



How much we actually get from the reaction

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

How much we predict we will get

Calculation of Atom Economy

$$\text{atom economy} = \frac{\text{mass of atoms in desired product}}{\text{mass of atoms in reactants}} \times 100\%$$

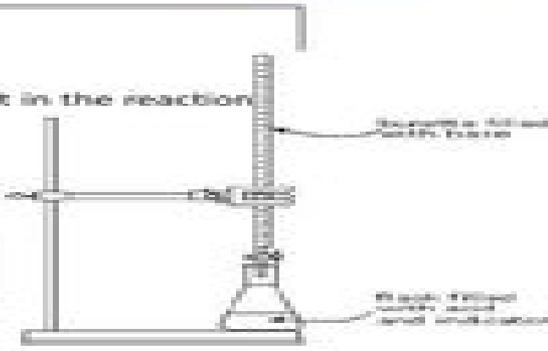
Titration calculations

- Step 1 - Write out the balanced symbol equation in the question
- Step 2 - Write all the information you have about each reactant/product in the reaction
- Step 3 - Make sure volumes are converted into **dm³**
- Step 4 - Calculate number of moles of the known concentration, then you will know moles of unknown.

You can then carry out concentration

Calculation of the unknown.

	HCl	NaOH	NaCl	H ₂ O
Amount				
Conc				
Volume				
Conc ³				
Volume				
Conc ²				
Conc				
Conc ¹				



Chemical Calculations

Sidney William Benson



Chemical Calculations:

Chemical Calculations Paul Yates, 1997-06-05 *Chemical Calculations* provides an introduction to the mathematics required for physical chemistry courses. This book is unique in that it provides a gentle introduction with a chemistry centered rather than math centered approach. Written by a chemist for undergraduate students it imparts an understanding of the subject from a chemist's viewpoint using examples from real chemistry. It includes illustrations that show exactly how to use calculators to work problems and examples of important chemical problems with fully worked solutions. This book is an ideal companion throughout a chemistry course that can be consulted when required and used to keep one step ahead of the lecture.

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