

ORGANOMETALLIC CHEMISTRY

Table: Common ligands and their electron contributions.

Ligand	Neutral atom	Oxidation State		Ligand	Neutral atom	Oxidation State	
		Electron contribution	Formal charge			Electron contribution	Formal charge
Carbonyl (M-CO)	2	2	0	Halogen (M-X)	1	2	-1
Phosphine (M-PR ₃)	2	2	0	Alkyl (M-R)	1	2	-1
Amine (M-NR ₂)	2	2	0	Aryl (M-Ar)	1	2	-1
Amide (M-NR ₂)	1	2	-1	Acyl (M-C(O)-R)	1	2	-1
Hydrogen (M-H)	1	2	-1	η^1 -Cyclopentadienyl	1	2	-1
Alkene (side-wise) η^2	2	2	0	η^1 -Allyl	1	2	-1
Alkyne (side-wise) η^2	2	2	0	η^2 -Allyl	3	4	-1
η^2 -C ₆₀	2	2	0	η^5 -Cyclopentadienyl	5	6	-1
Nitrosyl bent	1	2	-1	η^6 -Benzene	6	6	0
Nitrosyl Linear	3	2	+1	η^7 -Cycloheptatrienyl	7	6	+1
Carbene (M-CR ₂)	2	4	-2	Carbyne (M-CR)	3	6	-3
Alkoxide (M-OR)	1	2	-1	Thiolate (M-SR)	1	2	-1
μ -CO (M-CO-M)	2	2	0	μ -H	1	2	-1
μ -Alkyne	4	4	0	μ -X (M-X-M)	3	4	-1
μ -Alkyl	1	2	-1	X-Halogen			
μ -Phosphido (M-PR ₂ -M)	3	4	-1	μ -Amido (M-NR ₂ -M)	3	4	-1
				μ -Alkoxide (M-OR-M)	3	4	-1

Chemical Analysis Of Organometallic Comp

R Sandford



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