

Some Basic Principles of Chemistry

1. Suppose the elements X and Y combine to form two compounds XY_2 and X_3Y_2 . When 0.1 mole of XY_2 weighs 10 g and 0.05 mole of X_3Y_2 weighs 9 g, the atomic weights of X and Y are [NEET 2016 Phase II]
(a) 40, 30 (b) 60, 40 (c) 20, 30 (d) 30, 20
2. If Avogadro number N_A , is changed from $6.022 \times 10^{23} \text{ mol}^{-1}$ to $6.022 \times 10^{20} \text{ mol}^{-1}$ this would change [CBSE AIPMT 2015]
(a) the definition of mass in units of grams
(b) the mass of one mole of carbon
(c) the ratio of chemical species to each other in a balanced equation
(d) the ratio of elements to each other in a compound
3. 20.0 g of a magnesium carbonate sample decomposes on heating to give carbon dioxide and 8.0 g magnesium oxide. What will be the percentage purity of magnesium carbonate in the sample? (Atomic weight of Mg = 24) [CBSE AIPMT 2015]
(a) 75 (b) 96 (c) 60 (d) 84
4. What is the mass of precipitate formed when 50 mL of 16.9% solution of AgNO_3 is mixed with 50 mL of 5.8% NaCl solution? (Ag = 107.8, N = 14, O = 16, Na = 23, Cl = 35.5) [CBSE AIPMT 2015]
(a) 28 g (b) 3.5 g (c) 7 g (d) 14 g
5. The number of water molecules is maximum in [CBSE AIPMT 2015]
(a) 18 molecules water (b) 1.8 g of water
(c) 18 g of water (d) 18 moles of water
6. When 22.4 L of $\text{H}_2(\text{g})$ is mixed with 11.2 L of $\text{Cl}_2(\text{g})$, each at STP, the moles of $\text{HCl}(\text{g})$ formed is equal to [CBSE AIPMT 2014]
(a) 1 mole of $\text{HCl}(\text{g})$ (b) 2 moles of $\text{HCl}(\text{g})$
(c) 0.5 mole of $\text{HCl}(\text{g})$ (d) 1.5 moles of $\text{HCl}(\text{g})$
7. 1.0 g of magnesium is burnt with 0.56 g of oxygen in a closed vessel. Which reactant is left in excess and how much? [CBSE AIPMT 2014]
(At. weight of Mg = 24, O = 16)
(a) Mg, 0.16 g (b) O_2 , 0.16 g
(c) Mg, 0.44 g (d) O_2 , 0.28 g
8. How many grams of concentrated nitric acid solution should be used to prepare 250 mL of 2.0 M HNO_3 ? The concentrated acid is 70% HNO_3 . [NEET 2013]
(a) 45.0 g conc. HNO_3 (b) 90.0 g conc. HNO_3
(c) 70.0 g conc. HNO_3 (d) 54.0 g conc. HNO_3
9. 6.02×10^{20} molecules of urea are present in 100 mL of its solution. The concentration of solution is [NEET 2013]
(a) 0.02 M (b) 0.01 M
(c) 0.001 M (d) 0.1 M

Chem Principles Tbp

Robert R. Redfield



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Mendeleev Chemistry Journal, 1972 *Nuclear Science Abstracts*, 1975-06 **Process Chemistry** Francis Robert Bruce, 1956 **Clinical Chemistry** Michael L. Bishop, Janet L. Duben-Engelkirk, Edward P. Fody, 1985 **Fundamentals of Physical Chemistry for Students of Chemistry and Related Sciences** Arnold Eucken, 1925 *Nuclear Chemistry and Effects of Irradiation*, 1956 *Energy Research Abstracts*, 1985 Semiannual with semiannual and annual indexes

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