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# Voltas AC Remote Ultimate Guide

# Voltas Ac Remote Guide

**J Rink**



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web we then changed the volume between 5ml and 20ml and recorded the pressure at each this part of the lab showed that volume and pressure are inversely proportionate as the pressure would decrease if the volume was increased part b of the experiment looked at the relationship between temperature and pressure

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web procedure 1 the relationship between pressure and volume set the plunger of the syringe to 1 10 the capacity of the syringe the 2 0 ml mark if using a 20 ml syringe the volume is read from the edge of the bottom rib of the plunger attach the syringe to the connector at the end of sensor

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of

temperature rise that being 1/267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1/273 and concluded that the pressure would approach zero at 1/273 degrees c

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web the volume of 1.00 mol of any gas at stp standard temperature 273.15 k and pressure 1 atm is measured to be 22.414 l we can substitute 101.325 kpa for pressure 22.414 l for volume and 273.15 k for temperature into the ideal gas equation and solve for r

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web may 20 2018 french physicist jacques charles 1746 1823 studied the effect of temperature on the volume of a gas at constant pressure charles s law states that the volume of a given mass of gas varies directly with the absolute temperature of the gas when pressure is kept constant the absolute temperature is temperature measured

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273.15 k and 1 atm 101.325 kpa 1 at stp one mole of an ideal gas has a volume of about 22.4 l this is referred to as the standard molar volume figure 9 18

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts