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 $2 \text{ NaCl} + \text{ F}_2 \rightarrow 2 \text{ NaF} + \text{ Cl}_2$
 $4 \text{ H}_2 + \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$
 $5 \text{ Pb} + 2 \text{ HCl} \rightarrow 2 \text{ H}_2 + \text{ PbCl}_2$
 $6 \text{ Al} + 3 \text{ Br}_2 \rightarrow 2 \text{ AlBr}_3$
 $3 \text{ K}_2\text{SO}_4 + 6 \text{ KBr} + 1 \text{ Al}_2\text{SO}_4 + 3 \text{ H}_2\text{O}$
 $4 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
 $8 \text{ C} + 3 \text{ H}_2 + 5 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 4 \text{ H}_2\text{O}$

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 $2 \text{ C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $3 \text{ HBr} + \text{KHCO}_3 + \text{H}_2\text{O} \rightarrow \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SNO}_4$
 $4 \text{ NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3 \text{ NF}_3 + \text{SNF}_2 + \text{N}_2\text{O}_3$ using the equation from problem 2 above answer the following

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 $1 \text{ N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
 $2 \text{ C}_6\text{H}_{10} + 17 \text{ O}_2 \rightarrow 12 \text{ CO}_2 + 10 \text{ H}_2\text{O}$
 $3 \text{ HBr} + 1 \text{ KHCO}_3 + 1 \text{ H}_2\text{O} \rightarrow 1 \text{ KBr} + \text{CO}_2 + 2 \text{ GaBr}_3 + 3 \text{ Na}_2\text{SO}_3 + 1 \text{ Ga}_2\text{SO}_3 + 3 \text{ SO}_3 + 6 \text{ NaBr} + 3 \text{ SNO}_2 + 2 \text{ NF}_3 + 3 \text{ SNF}_2 + 1 \text{ N}_2\text{O}_3$ using the following equation
 $2 \text{ NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{ H}_2\text{O} + \text{Na}_2\text{SO}_4$

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