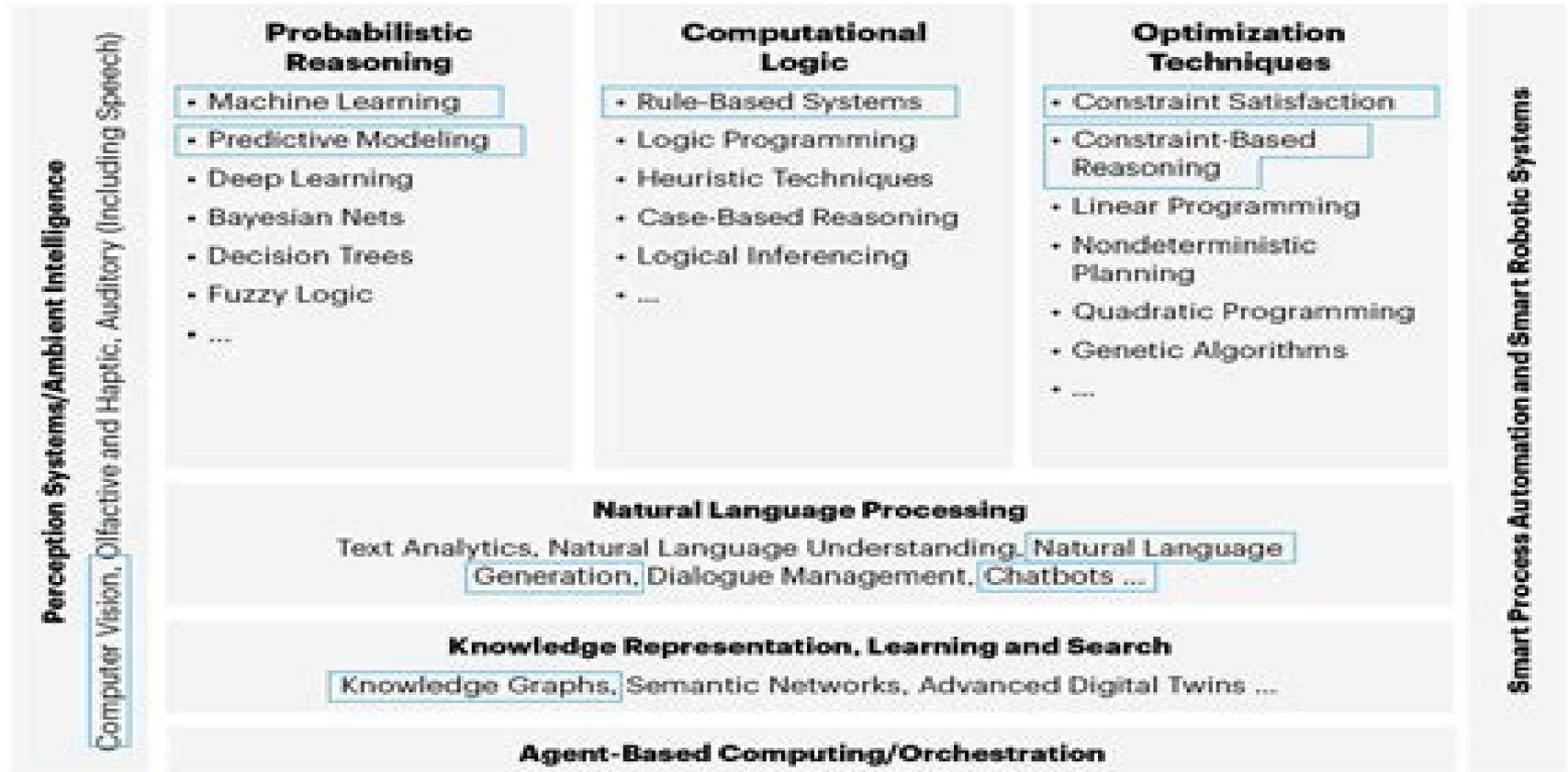


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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b Ba^{2+} c Br_2 d ca e Ga^{3+} f h 2 g h

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from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 $\text{Fe} + \text{Co} \rightarrow \text{Fe} + \text{Co} + 2\text{e}^-$ oxidation $\text{Co} + 2\text{e}^- \rightarrow \text{Co} + 2\text{e}^-$ reduction $\text{Fe} + 2\text{e}^- \rightarrow \text{Fe} + 2\text{e}^-$ 24 3 $\text{Ag} + \text{Ni} \rightarrow \text{Ag} + \text{Ni} + 3\text{e}^-$ oxidation $\text{Ni} + 3\text{e}^- \rightarrow \text{Ni} + 3\text{e}^-$ reduction 25 $\text{Cu} + \text{Pb} \rightarrow \text{Cu} + \text{Pb} + 2\text{e}^-$ oxidation $\text{Pb} + 2\text{e}^- \rightarrow \text{Pb} + 2\text{e}^-$ reduction 26 O_2

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