

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations from K_b and K_a values and vice versa from K_w values.
3. Calculate pH from pK_a (or pK_b).

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. proton indicator |
| 4. hydronium ion | 5. K_a value | 6. pH |
| 7. hydronium ion | 8. K_b value | |

Key Equations

- | | |
|---|---------------------------------------|
| 1. $K_w = [H_3O^+][OH^-] = 1.0 \times 10^{-14}$ | 2. $K_a = \frac{[H_3O^+][A^-]}{[HA]}$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pK_a = -\log K_a$ |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the reaction. Complete the reaction.

1. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
2. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
3. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
4. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
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8. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
9. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)
10. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (autoionization of water)

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19 2 Hydrogen Ions And Acidity Section Review Answers

LM Reder



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