

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH, and calculate the concentration of a weak acid from its pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---------------------------------------|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. acid ionization constant (K_a) | 6. hydronium ion concentration |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|------------------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (Eq. 17.1) | 2. $pH = -\log[H_3O^+]$ (Eq. 17.2) |
| 3. $pH = -\log[H_3O^+]$ (Eq. 17.3) | 4. $pH = -\log[H_3O^+]$ (Eq. 17.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step of the acid-base reaction shown in the completion reaction.

1. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (acidic solution)
2. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (acidic solution)
3. $H_2SO_4 + H_2O \rightleftharpoons HSO_4^- + H_3O^+$ (acidic solution)
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Didier Musso



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