

## 14.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous acids/bases.
2. Calculate hydroxide and hydronium ion concentrations and pH/pOH and the associated logarithmic functions.
3. Calculate the pH/pOH of an aqueous solution.

## Vocabulary

- |                   |                               |                  |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair   | 3. acid solution |
| 4. autoionization | 5. hydronium ion ( $H_3O^+$ ) | 6. hydroxide ion |
|                   | 7. acid solution              | 8. pH            |

## Key Equations

- |   |                         |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$                             | 4. $pOH = -\log[OH^-]$  |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Each correct response receives a score of one. Complete as needed.

1. What product is the \_\_\_\_\_ of a weak hydroxide ion and a hydroxide ion? \_\_\_\_\_ of the conjugate acid-base pair.
2. \_\_\_\_\_ of a weak acid and a hydroxide ion.
3. \_\_\_\_\_ of a weak acid and a hydroxide ion.
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20. \_\_\_\_\_ of a weak acid and a hydroxide ion.

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# 19 2 Hydrogen Ions And Acidity Section Review

**SJ Ball**



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