

14.5

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH and pOH and the associated logarithmic functions.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H_3O^+) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---|-------------------------|
| 1. $[H_3O^+][OH^-] = 1.0 \times 10^{-14}$ (at 25°C) | 2. $pH = -\log[H_3O^+]$ |
| 3. $pH = -\log[H_3O^+]$ | 4. $pOH = -\log[OH^-]$ |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid/base.

1. H_2O and H_3O^+ are _____, where H_2O is the _____.
2. H_2O and OH^- are _____, where H_2O is the _____.
3. H_2O and H_2O are _____, where H_2O is the _____.
4. H_2O and H_2O are _____, where H_2O is the _____.
5. H_2O and H_2O are _____, where H_2O is the _____.
6. H_2O and H_2O are _____, where H_2O is the _____.
7. H_2O and H_2O are _____, where H_2O is the _____.
8. H_2O and H_2O are _____, where H_2O is the _____.
9. H_2O and H_2O are _____, where H_2O is the _____.
10. H_2O and H_2O are _____, where H_2O is the _____.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

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Lauren Gardner



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