

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium | 5. hydroxide ion (OH^-) | 6. autoionization |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair indicated in the second step. Write the number in the space provided.

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) (14.1) (14.2) (14.3) (14.4)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
3. $\text{HCl} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{Cl}^-$ (14.1) (14.2) (14.3) (14.4)
4. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
5. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
6. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
7. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
8. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
9. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)
10. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.1) (14.2) (14.3) (14.4)

11. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

12. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

13. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

14. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

15. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

16. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

17. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

18. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

19. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

20. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

21. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

22. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$

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1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Work Sheet Answers Free and Paid eBooks

- 19 2 Hydrogen Ions And Acidity Work Sheet Answers Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Work Sheet Answers eBook Formats
- ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Work Sheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Work Sheet Answers
8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Work Sheet Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Work Sheet Answers
9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Work Sheet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Work Sheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Work Sheet Answers
- Setting Reading Goals 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
- Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Work Sheet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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