

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~Sulfuric acid~~ **H<sub>2</sub>SO<sub>4</sub>**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~ammonium hydroxide~~ **NH<sub>4</sub>OH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water

- HCl**      **HCl**      **HClO<sub>2</sub>**      **Mg(OH)<sub>2</sub>**      **HNO<sub>3</sub>**      **HNO<sub>2</sub>**      **HNO<sub>3</sub>**      **HNO<sub>2</sub>**

3. What is the formula for hydrosulfuric acid? **H<sub>2</sub>S**

4. **H<sub>2</sub>CO<sub>3</sub>** will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **H<sub>2</sub>CO<sub>3</sub> + H<sub>2</sub>O ⇌ H<sub>2</sub>O + HCO<sub>3</sub><sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>CO<sub>3</sub> is a weak acid.**
- Would you define **H<sub>2</sub>CO<sub>3</sub>** as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is **H<sub>2</sub>CO<sub>3</sub>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of **H<sub>2</sub>CO<sub>3</sub>**. **HCO<sub>3</sub><sup>-</sup>**
- Identify **H<sub>2</sub>CO<sub>3</sub>** as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. **F<sup>-</sup>** will react with water and gain a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **F<sup>-</sup> + H<sub>2</sub>O ⇌ HF + OH<sup>-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F<sup>-</sup> is a weak base.**
- Would you define **F<sup>-</sup>** as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H<sup>+</sup> ion.**
- In your written equation, is **F<sup>-</sup>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is gaining a proton.**
- Identify the conjugate acid of **F<sup>-</sup>**. **HF**
- Identify **F<sup>-</sup>** as polyprotic, amphoteric, or neither. **neither.**

6. **HCO<sub>3</sub><sup>-</sup>** will react with water and lose a H<sup>+</sup> ion in the process.

- Write a balanced equation for the reaction. **HCO<sub>3</sub><sup>-</sup> + H<sub>2</sub>O ⇌ H<sub>2</sub>O + CO<sub>3</sub><sup>2-</sup>**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HCO<sub>3</sub><sup>-</sup> is a weak acid.**
- Would you define **HCO<sub>3</sub><sup>-</sup>** as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
- In your written equation, is **HCO<sub>3</sub><sup>-</sup>** acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is losing a proton.**
- Identify the conjugate base of **HCO<sub>3</sub><sup>-</sup>**. **CO<sub>3</sub><sup>2-</sup>**
- Identify **HCO<sub>3</sub><sup>-</sup>** as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can lose the proton).**

# 19 3 Acids And Bases Answer Key Chemistry

**Michael Brown**



**19 3 Acids And Bases Answer Key Chemistry:**

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