



# Dog Like Jack

**RJ Shavelson**



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web a rate law shows how the rate of a chemical reaction depends on reactant concentration for a reaction such as  $aA \rightarrow bB$  products the rate law generally has the form  $\text{rate} = k[A]^n$  where  $k$  is a proportionality constant called the rate constant and  $n$  is the order of the reaction with respect to  $A$  the value of  $n$  is not related to the reaction stoichiometry and

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web the rate determining slower step gives a rate law showing second order dependence on the  $[A]^2$  concentration and the sum of the two elementary equations gives the overall net reaction in general when the rate determining slower step is the first step in the reaction mechanism the rate law for the overall reaction is the same as the rate

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web rate laws and reaction order the relation between the rate of a reaction and the concentrations of reactants is expressed by its rate law for example the rate of the gas phase decomposition of dinitrogen pentoxide  $2N_2O_5 \rightarrow 4NO_2 + O_2$  has been found to be directly proportional to the concentration of  $N_2O_5$  text rate  $k[A]^n$

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web the rate determining step limits the overall rate and therefore determines the rate law for the overall reaction introduction multistep reactions one of the most important applications of kinetics is to the study of reaction mechanisms or the sequences of steps by which chemical reactions occur

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web the initial rate law is then measured for each of the reactions consider the reaction between nitrogen monoxide gas and hydrogen gas to form nitrogen gas and water vapor the following data were collected for this reaction at 1280 text o text c see table below notice that the starting concentrations of ce no and ce h

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web determination of a rate law and activation energy in this experiment glass beakers pipettes and a hotplate were utilized to set up the procedures a solution of deionized water a buffer potassium iodide ki sodium thiosulfate na<sub>2</sub>s<sub>2</sub>o<sub>3</sub> starch and hydrogen peroxide h<sub>2</sub>o<sub>2</sub> were used in order to calculate the rate of reaction

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web determining the rate law page 1 page 2 page 3 kinetics experiments the goal of a kinetics experiment is to measure the concentration of a species at a particular time during a reaction so that a rate law can be determined

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web core concepts some reactions will go fast and some will go slow the speed of the reaction is it s reaction rate which is dictated by a rate law in this article we will learn about reaction rates rate laws the rate constant and the reaction order

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the first is a problem of applied thermodynamics the second is the central problem of chemical kinetics

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