

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n = moles and R = 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

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Directory of Distance Learning Opportunities Modoc Press, Inc., 2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154 institutions or distance learning consortium members Following an introduction that describes existing practices and delivery methods the Directory offers three indexes Subject Index of Courses Offered by Level Course Level Index Geographic Index All information was supplied by the institutions Entries include current contact information a description of the institution and the courses offered grade level and admission information tuition and fee information enrollment periods delivery information equipment requirements credit and grading information library services and accreditation *United States Air Force Academy* United States Air Force Academy, 1983 EC Study Guide to Environment-related Courses Institute for European Environmental Policy, Commission of the European Communities. Directorate-General Environment, Nuclear Safety and Civil Protection, 1993 Annual Catalogue United States Air Force Academy, 1984 *Instructor's manual and solutions to problems to accompany General chemistry: principles and structures* James E. Brady, 1975-04 **El-Hi Textbooks in Print**, 1977 *How to Pass Annapolis Entrance Exams* Arco Publishing Company, 1959 **Cliffs TASP Preparation Guide** Jerry Bobrow, 2007-08-13 The CliffsTestPrep series offers full length practice exams that simulate the real tests proven test taking strategies to increase your chances at doing well and thorough review exercises to help fill in any knowledge gaps Cliffs TASP Preparation Guide can help you get ready for the Texas Academic Skills Program Since the TASP requires you to use some basic skills you may not have used recently thorough preparation is the key to doing your best This guide developed by test preparation experts and instructors contains materials techniques and strategies for taking the TASP that have been carefully researched and tested and are currently used in college and teachers association preparation programs throughout the country In this guide you ll find Two full length practice tests An overview of the different sections of the test Sample questions and English review and a writing sample Answers and complete explanations of all questions on the practice tests The TASP measures your abilities in three areas reading mathematics and writing This book will help you understand the different types of questions that appear in each section of the test giving you clear explanations of the directions as well as plenty of sample questions to help sharpen your test taking skills With guidance from the CliffsTestPrep series you ll feel at home in any standardized test environment **The National Guide to Educational Credit for Training Programs** American Council on Education, 2005 Highlights over 6 000 educational programs offered by business labor unions schools training suppliers professional and voluntary associations and government agencies **Guide to American Graduate Schools** Harold Doughty, 1997 Anyone faced with the overwhelming selection of graduate schools knows how valuable clear and concise information can be And that s just what they ll get with this comprehensive guide now fully updated and in its

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Peterson's Guide to Graduate Programs in the Biological Sciences 1997 Peterson's,1997-01-05

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Treasury Enforcement Agent Arco Publishing Company,1961

Geology Edward C.

Gruber,1963

How to Pass High on the U.S. Coast Guard Academy Tests Arco Publishing Company,1959

Subject

Guide to Books in Print ,1975

Using Physical Models of Biomolecules to Teach Concepts of Biochemical

Structure in Introductory Undergraduate Chemistry John Yi,2004

Paperbound Books in Print ,1992

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