

## Chemistry 338 2<sup>nd</sup> Semester Final Exam Review Sheet

### Unit 1 - Matter & Measurement – Chapters 1 & 2

1. A student experimentally determines the density of an object to be 4.2 g/mL. If the accepted value is 5.0 g/mL, calculate the % error.

$$\% \text{ error} = \frac{\text{experimental} - \text{accepted}}{\text{accepted}} \times 100$$

$$\% \text{ error} = \frac{4.2 - 5.0}{5.0} \times 100 = -16\% \text{ error}$$

### Unit 2 – Periodic Table – Chapter 5

Element	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #	Charge
Tungsten ion	$^{184}_{74}\text{W}^{+2}$	74	74	110	72	184	+2

3. Using the following atomic radius values, predict the atomic radius of potassium (K):  
Li – 145 pm, Na – 180 pm, K - ?, Rb – 235 pm  
K – 220 pm

### Unit 3 – Reactions/Compounds – Chapters 7 & 8

4. Classify the compound as ionic or covalent. Name the compound
- $\text{Fe}_2\text{O}_3$  - iron (III) oxide (ionic - roman numerals for transition metals, change ending to -ide)
  - $\text{Cl}_2\text{O}_8$  - dichlorine octaoxide (covalent - use prefixes, change ending to -ide)
5. Classify the compound as ionic or covalent. Write the formula
- Trisulfur Dinitride –  $\text{S}_3\text{N}_2$  (covalent)
  - Zinc Hydroxide –  $\text{Zn}(\text{OH})_2$  (ionic - look up ion charges and criss-cross)  $\text{Zn}^{+2} \text{OH}^{-1}$
  - Chromium (III) nitrate –  $\text{Cr}(\text{NO}_3)_3$  (ionic - criss-cross)  $\text{Cr}^{+3} \text{NO}_3^{-2}$
6. Write the equation, balance it, and classify the type of reaction:  
Iron (II) sulfate reacts with sodium hydroxide to yield iron(II) hydroxide and sodium sulfate
- $$1 \text{ FeSO}_4 + 2 \text{ NaOH} \rightarrow 1 \text{ Fe}(\text{OH})_2 + 1 \text{ Na}_2\text{SO}_4$$
- double displacement (two ionic compounds reacting)
7. Balance the equation and classify the type of reaction:
- $$1 \text{ Cu} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Cu}(\text{NO}_3)_2 + 2 \text{ Ag}$$
- Single displacement (element and a compound reacting)

### Unit 4 – Moles – Chapter 3

8. Calculate the molar mass of  $\text{Fe}_2(\text{SO}_4)_3$ .
- $$\begin{aligned} \text{Fe } (2)(55.9 \text{ g/mol}) &= 111.8 \text{ g/mol} \\ \text{S } (3)(32.1 \text{ g/mol}) &= 96.3 \text{ g/mol} \\ + \text{O } (12)(16.0 \text{ g/mol}) &= +192 \text{ g/mol} \\ &= 400.1 \text{ g/mol} \end{aligned}$$

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