

# Atomic Structure (Formula)

1.  $E = -13.6 \frac{Z^2}{n^2} \text{ eV}$

2.  $v = 2.18 \times 10^6 \frac{Z}{n} \text{ m/sec}$

3.  $r = 0.529 \frac{n^2}{Z} \text{ \AA}$

4.  $E(n) = -\frac{R_H}{n^2}$   $R_H = 2.18 \times 10^{-18}$

5.  $\bar{\nu} = \frac{1}{\lambda} = R_H Z^2 \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$  (for  $E_{\text{photon}} = h\nu = hc$ )

7.  $\lambda = \frac{h}{mv} = \frac{h}{\sqrt{2mKE}} = \frac{h}{\sqrt{2m_e V}}$   $\Delta V_{SP} = \sqrt{V_e - V_{SP}} = KE$

9. Photoelectric effect

$$h\nu = W_0 + \frac{1}{2}mv_e^2$$

10. Heisenberg

i)  $\Delta x \cdot \Delta p \geq \frac{h}{4\pi}$

11.  $mvr = \frac{nh}{2\pi}$

↑  
orbit ang. mom

ii)  $\Delta x \cdot \Delta v \geq \frac{h}{4\pi m}$

iii)  $\Delta E \cdot \Delta t \geq \frac{h}{4\pi}$

12. orbital ang. mom =  $\sqrt{l(l+1)} \frac{h}{2\pi}$

13. total spin (S) =  $\frac{1}{2} \times n \rightarrow$  no. of unpaired e<sup>-</sup>

14. spin magnetic moment =  $\sqrt{n(n+2)} \text{ B.M.}$

15. Spin multiplicity =  $2S+1$

16. total nodes =  $n-1$

17. radial nodes =  $n-l-1$

18. ang. nodes (N.P) =  $l$

19. max type of wavelength =  $\frac{n(n-1)}{2}$

20. total / max electron =  $2n^2$   
↳ orbit no.

# Calculation Of Atomic Structures

**SB Merriam**



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