

Chemical Bonding**SECTION 3**

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 - formula unit.
 - molecule.
 - crystal.
 - atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 - molecules.
 - positive ions.
 - dipoles.
 - negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 - higher in potential energy.
 - lower in potential energy.
 - equal in potential energy.
 - unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 - Compound A is not an ionic compound.
 - It will be more difficult to break the bonds in compound A than those in compound B.
 - Compound B has larger crystals than compound A.
 - Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 - stronger than the attractive forces among formula units in ionic bonding.
 - weaker than the attractive forces among formula units in ionic bonding.
 - approximately equal to the attractive forces among formula units in ionic bonding.
 - equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

Section 3 1 Chemistry Answers

Roman Wölfel



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