

9.1

THE ARITHMETIC OF EQUATIONS

SECTION REVIEW

Objectives

- Calculate the amount of reactants required or product formed in a stoichiometric process.
- Interpret balanced chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

Key Terms

- stoichiometry

Part A Completion

Copy this completion exercise on blank paper and demonstrate your understanding of the concepts and terms in this section by completing the sentences. Write down the complete word or phrase, short phrase, or name first.

- The coefficients in a balanced chemical equation indicate _____ 1. _____ the relative number of _____ 2. _____ of reactants and products. All stoichiometric calculations begin with a _____ 3. _____ (molar _____ 4. _____) and _____ 5. _____ are used to find the mass, moles, volume, and representative particles that react.
- In solving stoichiometric problems, conversion factors relating moles of reactants to _____ 6. _____ of products are used.
- If you assume _____ 7. _____, the equation also tells you about the volume of gas.

Part B True-False

Circle the word *if* in the statements to indicate they are (C) correct or (I) incorrect. Write (C) or (I) on the line.

- _____ 7. The coefficients in a balanced chemical equation can be used to find the relative molar masses of reacting reactants and products.
- _____ 8. The coefficients in a balanced chemical equation tell the relative numbers of reactants and products, expressed in any suitable units of volume.
- _____ 9. To calculate the mass of a molecule in grams, you can use the molar mass and Avogadro's number.

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