

## 16.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                                    |                   |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair        | 3. weak acid      |
| 4. hydronium      | 5. hydroxide ion ( $\text{OH}^-$ ) | 6. autoionization |
|                   | 7. weak base                       | 8. pH             |

## Key Equations

- |  |  |
|--|--|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$ |
| 3. $\text{pH} = -\log [\text{H}_3\text{O}^+]$  | 4. $\text{pH} = \text{pK}_a + \log \frac{[\text{A}^-]}{[\text{HA}]}$                             |

## Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify hydroxide ion as the conjugate base of the hydronium ion and vice versa)
2.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
3.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
7.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)
8.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (identify water as the conjugate acid of the hydroxide ion)

1. **conjugate acid**

2. **conjugate base**

3. **acid**

4. **conjugate acid**

5. **acid**

6. **conjugate acid**

7. **conjugate acid**

8. **acid**

9. **conjugate acid**

10. **conjugate acid**

11. **conjugate acid**

# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Shasha Hu**



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