

Exercise 1

For (Part 'B')

- If a 0.1 M solution of HCN is 0.01% ionised, the ionisation constant for HCN is
(a) 10^{-6} (b) 10^{-2} (c) 10^{-8} (d) 10^{-5}
- A 0.1 N solution of sodium bicarbonate has a pH value of
(a) 8.4 (b) 4.8 (c) 4.0 (d) 13
- Which of the following is not a Lewis acid?
(a) AlCl_3 (b) CO_2 (c) SO_2 (d) CN^-
- The hydrogen ion concentration of a slightly acidic water solution can be represented by
(a) $10^{-\text{pOH}}$ (b) $14 - \text{pOH}$
(c) $\frac{K_w}{\text{pOH}}$ (d) $10^{-(14 - \text{pOH})}$
- According to the protonic concept of acids and bases, which one of the following is a substance classified as weak acid?
(a) NH_4^+ (b) NO_3^- (c) ClO_4^- (d) SO_4^{2-}
- The conjugate base of NH_4^+ is
(a) NH_4OH (b) NH_3 (c) OH^- (d) NH_2
- According to Bronsted definition of acids and bases
(a) a base may be an electrically neutral substance
(b) the same substance cannot function both as acid and base
(c) any acid can yield a base by gaining a proton
(d) a base cannot be a cation
- Of the following ions the one which is an ampholyte is
(a) SO_4^{2-} (b) NO_3^- (c) PO_4^{3-} (d) HPO_4^{2-}
- Which one of the following species may be Lewis base?
(a) Negative ions
(b) Positive ions
(c) Inorganic molecules with double bond
(d) Organic molecules with sigma bond
- The conjugate base of H_2O^+ is
(a) OH^- (b) H_2O (c) H_2O_2 (d) H^+
- The most basic species among the following
(a) CH_3^- (b) H_2O (c) F^- (d) OH^-
- According to Pearson, a hard base is one whose donor atom has
(a) high electronegativity, high polarisability and easy to oxidise
(b) low electronegativity, low polarisability and which is difficult to oxidise
(c) high electronegativity, low polarisability and which is difficult to oxidise
(d) low electronegativity, high polarisability and difficult to oxidise
- The salt of which one of the following five weak acids will be the most hydrolysed?
(a) $\text{HD} : K_a = 4 \times 10^{-10}$ (b) $\text{HC} : K_a = 3 \times 10^{-8}$
(c) $\text{HB} : K_a = 2 \times 10^{-5}$ (d) $\text{HA} : K_a = 1 \times 10^{-4}$
- Two hypothetical acids HA and HB have the dissociation constant 1×10^{-3} and 1×10^{-5} respectively in water. How many times HA is stronger than HB?
(a) 1000 times (b) 100 times
(c) 10 times (d) Not definite
- Which of the following is a hard acid?
(a) Ag^+ (b) Au^+ (c) Li^+ (d) Cu^+
- Which of the following is a soft base?
(a) H^- (b) CO_3^{2-}
(c) NO_3^- (d) CH_3COO^-
- Which is the correct order of decreasing basicity of halide ion?
(a) $\text{F}^- > \text{Cl}^- > \text{Br}^- > \text{I}^-$ (b) $\text{Cl}^- > \text{Br}^- > \text{I}^- > \text{F}^-$
(c) $\text{F}^- > \text{I}^- > \text{Br}^- > \text{Cl}^-$ (d) $\text{I}^- > \text{Br}^- > \text{Cl}^- > \text{F}^-$
- Which of the following is the most acidic in water?
(a) HF (b) NH_3 (c) CH_4 (d) H_2O
- Which of the following cationic species can act as Bronsted-Lowry base?
(a) $[\text{Al}(\text{H}_2\text{O})_6]^{3+}$ (b) $[\text{Fe}(\text{H}_2\text{O})_5(\text{OH})]^{2+}$
(c) NH_4^+ (d) $[\text{Fe}(\text{H}_2\text{O})_6]^{3+}$
- Which of the following is not a Lewis acid?
(a) SbF_5 (b) SO_2 (c) AlCl_3 (d) CN^-
- Which of the following statement is true?
(a) Hard acids are non-polarizable
(b) Pyridine, SO_3 and H_2O are Lewis bases
(c) According to HSAB principle, hard acid prefer to coordinate with soft base
(d) According to solvent system concept of acids and bases, NH_4Cl behaves as a base in liquid ammonia
- Which of the following will act as an acid in SO_2 according to solvent system concept?
(a) SOCl_2 (b) Na_2SO_3
(c) K_2SO_3 (d) HCl
- Which of the following is not a hard base according to Pearson's concept of hard and soft bases?
(a) O^{2-} (b) NO_3^-
(c) I^- (d) F^-

Context Questions For Advanced Chemistry

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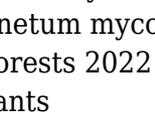
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