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**Kyle Richter**



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web a model of acids and bases in which an acid is a hydrogen ion donor and a base is a hydrogen ion acceptor conjugate

acid the species produced when a base accepts a hydrogen ion from an acid

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web arrhenius model of acids a substance that contains hydrogen and ionizes to produce hydrogen ion in aqueous solution

arrhenius model of bases a substance that contains a hydroxide group and dissociates to produce a hydroxide

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web kw the product of the concentrations of the hydrogen ions and hydroxide ions acidic solution one which  $\text{H}^+$  is greater than

the  $\text{OH}^-$  basic solution one in which the  $\text{H}^+$  is less than the  $\text{OH}^-$  alkaline solutions the  $\text{H}^+$  of a basic solution is less than  $1 \times 10^{-7} \text{ M}$

they are the same as basic solutions

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web jan 10 2021 to identify the strongest base we can determine their weakest conjugate acid the conjugate acids of  $\text{CH}_3\text{NH}_2$  and  $\text{S}^{2-}$  are  $\text{CH}_4$   $\text{NH}_3$  and  $\text{HS}^-$  respectively next we consider that acidity increases with positive charge on the molecule

thus ruling out that  $\text{S}^{2-}$  is the weakest base

they are the same as basic solutions

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web compounds formed from combinations of acids and bases taste salty and are referred to in chemistry as salts sweet

compounds have characteristics of both acids and bases in the same molecule we will explore the relationship between

molecular structure and acids bases and consider water solutions of acids and bases

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web conjugate acid base pair two substances related by the loss or gain of a single hydrogen ion hydronium ion  $\text{H}_3\text{O}^+$  a water

molecule that gains a hydrogen ion becomes this which is positively charged what did lewis propose an acid accepts a pair of

electrons during a reaction while a base donates a pair of electrons

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web solutions manual acids i and bases section 18 1 introduction to acids and bases pages 634 643 practice problems pages 635 640 write balanced equations for reactions between the following aluminum and sulfuric acid  $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow 2\text{Al}^{3+} + 3\text{H}_2 + 3\text{SO}_4^{2-}$  calcium carbonate and hydrobromic acid