

## 16.1

## HYDROGEN IONS AND ACIDITY

## Section Overview

## Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

## Vocabulary

- |                   |                                    |                   |
|-------------------|------------------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair        | 3. weak acid      |
| 4. hydronium      | 5. hydroxide ion ( $\text{OH}^-$ ) | 6. autoionization |
|                   | 7. weak base                       | 8. pH             |

## Key Equations

- |                                                                                                    |                                                                                                  |
|----------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------|
| 1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$ |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$                                                       | 4. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$                 |

## Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$  (strong hydroxide base) / autoionization of  $\text{H}_2\text{O}$  / the autoionization of  $\text{H}_2\text{O}$  which produces an acidic solution / 2 / none
2.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$  (strong acid) / 2 / autoionization of  $\text{H}_2\text{O}$  / the acidic solution / 2 /  $\text{H}_2\text{SO}_4$  / 1
3.  $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightleftharpoons \text{HSO}_4^- + \text{H}_3\text{O}^+$  (strong acid) / 2 / autoionization of  $\text{H}_2\text{O}$  / the acidic solution / 2
4.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$  (strong hydroxide base) / autoionization of  $\text{H}_2\text{O}$  / the basic solution / 2 /  $\text{OH}^-$  / 1
5.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$  (strong hydroxide base) / autoionization of  $\text{H}_2\text{O}$  / the basic solution / 2
6.  $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{O} + \text{OH}^-$  (strong hydroxide base) / autoionization of  $\text{H}_2\text{O}$  / the basic solution / 2

1. autoionization
2.  $\text{H}_2\text{SO}_4$  / 1
3.  $\text{H}_2\text{SO}_4$  / 1
4. autoionization of  $\text{H}_2\text{O}$
5. autoionization of  $\text{H}_2\text{O}$
6. autoionization of  $\text{H}_2\text{O}$
7.  $\text{OH}^-$  / 1
8. autoionization of  $\text{H}_2\text{O}$
9. autoionization of  $\text{H}_2\text{O}$
10. autoionization of  $\text{H}_2\text{O}$

# 19 2 Hydrogen Ions And Acidity Answer Key

**Rachel Sandford**



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