

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, molar ratios, and pH values, and vice versa, for weak acid solutions (at 25°C).
3. Calculate the pH of weak (or strong) acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. hydroxide ion | 5. weak acid | 6. autoionization |
| | 7. weak solution | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{pH} = -\log_{10} [\text{H}^+]$ (Eq. 14.1) | 2. $\text{pOH} = -\log_{10} [\text{OH}^-]$ (Eq. 14.2) |
| 3. $\text{pH} + \text{pOH} = \text{pK}_w$ (Eq. 14.3) | 4. $\text{pH} + \text{pOH} = 14.00$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number to indicate the correctness of the response to each item. (0 = incorrect; 1 = correct; 2 = partially correct; 3 = correct; 4 = completely correct.)

1. When a substance is _____, it is a weak hydroxide solution.
2. _____ is the concentration of ions which produces an acidic solution. _____ is _____.
3. _____ is the concentration of _____.
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19 2 Hydrogen Ions And Acidity Section Review Answers

Anthony S. Fauci



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