

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of aqueous concentrations.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa from [OH⁻] and [H⁺].
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion (H ₃ O ⁺) | 6. autoionization |
| | 7. acid solution | 8. pH |

Key Equations

- | | |
|---------------------------------------|---------------------------------------|
| 1. $[H^+] = 10^{-pH}$ (Equation 14.1) | 2. $pH = -\log [H^+]$ (Equation 14.2) |
| 3. $pH = -\log [H^+]$ (Equation 14.3) | 4. $[H^+] = 10^{-pH}$ (Equation 14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the conjugate acid-base pair represented in the reaction. Also indicate which member of the pair is the stronger acid.

1. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (strong hydroxide base) / autoionization of H_2O / the autoionization of H_2O which produces an acidic solution / 2 / weak
2. $H_2PO_4^- + H_2O \rightleftharpoons HPO_4^{2-} + H_3O^+$ / acid / 1 / weak
3. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (strong hydroxide base) / autoionization of H_2O / the autoionization of H_2O which produces an acidic solution / 2 / weak
4. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (strong hydroxide base) / autoionization of H_2O / the autoionization of H_2O which produces an acidic solution / 2 / weak
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7. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (strong hydroxide base) / autoionization of H_2O / the autoionization of H_2O which produces an acidic solution / 2 / weak
8. $H_2O + H_2O \rightleftharpoons H_3O^+ + OH^-$ (strong hydroxide base) / autoionization of H_2O / the autoionization of H_2O which produces an acidic solution / 2 / weak

1. **conjugate acid**

2. **conjugate base**

3. **acid**

4. **stronger acid**

5. **stronger acid**

6. **stronger acid**

7. **stronger acid**

8. **stronger acid**

9. **stronger acid**

10. **stronger acid**

11. **stronger acid**

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Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings

- 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
- 5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
- 6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key
- 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key
- 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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