

17.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH, and calculate hydroxide ion concentration from pH.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|--------------------------------|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. proton indicator |
| 4. hydronium ion | 5. hydronium ion concentration | 6. hydronium ion concentration |
| | 7. hydroxide ion | 8. pH |

Key Equations

- | | |
|---|--|
| 1. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 17.1) | 2. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ (Eq. 17.16) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (Eq. 17.2) | 4. $\text{pH} = \text{pK}_a + \log\left(\frac{[\text{A}^-]}{[\text{HA}]}\right)$ (Eq. 17.16) |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)
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10. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (autoionization of water)

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19 2 Hydrogen Ions And Acidity Section Review

Lauren Gardner



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