

Year 1 Required Inorganic Equations

Name	Equations
1) Reactions of Mg - Ba with water	$\textcircled{1} \text{ Mg}_{(s)} + 2\text{H}_2\text{O}_{(l)} \longrightarrow \text{Mg}(\text{OH})_2_{(aq)} + \text{H}_2_{(g)}$ <p style="text-align: center; color: red;">OR</p> $\text{Mg}_{(s)} + \text{H}_2\text{O}_{(g)} \longrightarrow \text{MgO}_{(s)} + \text{H}_2_{(g)}$ $\textcircled{2} \text{ Ca}_{(s)} + 2\text{H}_2\text{O}_{(l)} \longrightarrow \text{Ca}(\text{OH})_2_{(aq)} + \text{H}_2_{(g)}$ $\textcircled{3} \text{ Sr}_{(s)} + 2\text{H}_2\text{O}_{(l)} \longrightarrow \text{Sr}(\text{OH})_2_{(aq)} + \text{H}_2_{(g)}$ $\textcircled{4} \text{ Ba}_{(s)} + 2\text{H}_2\text{O}_{(l)} \longrightarrow \text{Ba}(\text{OH})_2_{(aq)} + \text{H}_2_{(g)}$
2) Extraction of Titanium	$\text{TiCl}_4_{(l)} + 2\text{Mg}_{(s)} \longrightarrow \text{Ti}_{(s)} + 2\text{MgCl}_2_{(s)}$
3) Use of CaO/Ca(OH) ₂ to remove SO ₂ from the gas	$\textcircled{1} \text{ CaO}_{(s)} + \text{SO}_2_{(g)} \longrightarrow \text{CaSO}_3_{(s)}$ $\textcircled{2} \text{ CaCO}_3_{(s)} + \text{SO}_2_{(g)} \longrightarrow \text{CaSO}_3_{(s)} + \text{CO}_2_{(g)}$
4) Reactions of Sodium Halides + H ₂ SO ₄ <small>(Can be worked out)</small>	<p style="color: red;">① NaCl / NaBr / NaI:</p> $\text{NaCl}_{(s)} + \text{H}_2\text{SO}_4_{(l)} \longrightarrow \text{NaHSO}_4_{(s)} + \text{HCl}_{(g)}$ <p style="color: red;">② NaBr:</p> $\text{H}_2\text{SO}_4_{(l)} + 2\text{H}^+ + 2\text{Br}^- \longrightarrow \text{Br}_2_{(g)} + \text{SO}_2_{(g)} + 2\text{H}_2\text{O}_{(l)}$ <p style="color: red;">NaI:</p> $\text{H}_2\text{SO}_4_{(l)} + 2\text{H}^+ + 2\text{I}^- \longrightarrow \text{I}_2_{(s)} + \text{SO}_2_{(g)} + 2\text{H}_2\text{O}_{(l)}$ <p style="color: red;">③ NaI:</p> $\text{H}_2\text{SO}_4_{(l)} + 6\text{H}^+ + 6\text{I}^- \longrightarrow 3\text{I}_2_{(s)} + \text{S}_{(s)} + 4\text{H}_2\text{O}_{(l)}$ <p style="color: red;">④ NaI:</p> $\text{H}_2\text{SO}_4_{(l)} + 8\text{H}^+ + 8\text{I}^- \longrightarrow 4\text{I}_2_{(s)} + \text{H}_2\text{S}_{(g)} + 4\text{H}_2\text{O}_{(l)}$
5) Formation of Bleach	$2\text{NaOH}_{(aq)} + \text{Cl}_2_{(g)} \longrightarrow \text{NaClO}_{(aq)} + \text{NaCl}_{(aq)} + \text{H}_2\text{O}_{(l)}$
6) Chlorine + Water <small>(normal conditions)</small>	$\text{Cl}_2_{(g)} + \text{H}_2\text{O}_{(l)} \rightleftharpoons \text{HClO}_{(aq)} + \text{HCl}_{(aq)}$

A Level Guide Chemistry

Camilla Rothe



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