

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO₃~~ **nitric acid**
- ~~HNO₃~~ **nitric acid**
- ~~Sulfuric acid~~ **H₂SO₄**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)₂**
- ~~ammonium hydroxide~~ **NH₄OH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

- HCl** ~~HCl~~ ~~HCl~~ ~~Mg(OH)₂~~ **HNO₃** **H₂SO₄** ~~HNO₂~~ **HClO₄**

3. What is the formula for hydronium? **H₃O⁺**

4. H_2CO_3 will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{H}_2\text{CO}_3 + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{HCO}_3^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂CO₃ is a weak acid.**
- Would you define H_2CO_3 as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is H_2CO_3 acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of H_2CO_3 . **HCO_3^-**
- Identify H_2CO_3 as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F^- will react with water and gain a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{F}^- + \text{H}_2\text{O} \rightleftharpoons \text{HF} + \text{OH}^-$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F^- is a weak base.**
- Would you define F^- as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H^+ ion.**
- In your written equation, is F^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
- Identify the conjugate acid of F^- . **HF**
- Identify F^- as polyprotic, amphoteric, or neither. **neither**

6. HCO_3^- will react with water and lose a H^+ ion in the process.

- Write a balanced equation for the reaction. **$\text{HCO}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{CO}_3 + \text{CO}_3^{2-}$**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HCO_3^- is a weak acid.**
- Would you define HCO_3^- as a proton donor or a proton acceptor? **proton donor, because it is losing a H^+ ion.**
- In your written equation, is HCO_3^- acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of HCO_3^- . **CO_3^{2-}**
- Identify HCO_3^- as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can act like an acid).**

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