

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide and hydronium ion concentrations and pH values and vice versa for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|---|--------------------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid |
| 4. hydronium ion | 5. acid dissociation constant (K_a) | 6. hydronium ion concentration |
| | 7. weak base | 8. pH |

Key Equations

- | | |
|--|--|
| 1. $K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$ (Eq. 14.1) | 2. $K_b = \frac{[\text{B}][\text{H}_2\text{O}]}{[\text{BH}^+]}$ (Eq. 14.2) |
| 3. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (Eq. 14.3) | 4. $\text{pOH} = -\log[\text{OH}^-]$ (Eq. 14.4) |

Part A: Completion

Use the appropriate number of significant figures in your answers. If the answer requires that you round to the nearest integer, indicate whether you are rounding up or down.

1. What products are _____ when hydrochloric acid and methanol react? _____ the conjugate base of the acid.
2. _____ is the conjugate base of _____.
3. _____ is the conjugate acid of _____.
4. _____ is the conjugate base of _____.
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20. _____ is the conjugate base of _____.

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19 2 Hydrogen Ions And Acidity Work Sheet Answers

W. M. Clark



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