

Distributive Property. Color By Code.



Solve the math problems. Match the answer to a letter from the answer box. Color the picture by the letter.

Answer Box

A	B	C	D	E	F	G	H	I	J
$10x + 25$	$-40x - 36$	$45x - 10$	$2x - 3$	$3x - 5$	$-30x - 45$	$12x - 30$	$14x - 12$	$-8x - 6$	$-40x - 6$

No.		Letter	Color
1p	$1(2x - 5) =$		Orange
2p	$1(2x - 3) =$		Dark Blue
3p	$6(2x - 5) =$		Brown
4p	$2(7x - 8) =$		Yellow
5p	$2(5x - 1) =$		Light Grey

No.		Letter	Color
6p	$-4(2x + 2) =$		Brown
7p	$5(2x + 5) =$		Dark Blue
8p	$-3(8x + 9) =$		Yellow
9p	$9(5x - 4) =$		Orange
10p	$5(5x - 2) =$		Light Grey

Distributive Property Coloring

**Arthur F. Coxford, James T.
Fey, Christian R. Hirsch, Harold L.
Schoen, Eric W. Hart, Brian A.
Keller, Ann E. Watkins, Rebecca K.
Walker**

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web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = 44.2 \text{ kJ mol}^{-1}$ write equations to similarly represent the dissolving process for

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web hand warmers are usually placed inside the gloves to keep your hands warm the heat keeping your hands warm is formed as a result of an exothermic reaction which occurs in a small packet hand warmers have 5 main ingredients in them salt carbon polypropylene vermiculite water and iron

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution ΔH_{soln} or the amount of energy released or absorbed in kJ mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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