

**RICK TOWNSEND**

209'1" (63.7m)  
239'9" (73.08m)

84'11" (25.88m)  
102'5" (31.22m)

36"  
(10.97 m)

# Boeing 777 Study Guide 2023 Edition

Nose Radius  
110' (33.5m)  
131.2' (40.0 m)

Tail Radius  
95' (29.0 m)  
147.1' (44.8m)

Wing Tip Radius  
145' (44.2m)  
160.2' (48.8m)

**Covering the  
777-200 & 777-300  
Versions**

... for minimum  
turning radius. (Slow continuous  
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# Boeing 777 Study Guide

**Jessica J Manson**



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mole moles of  $\text{H}_2\text{O}$  40.1822 moles therefore according to the equation

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 time the total pressure  $p_a \times \text{total mole fraction}$  a left frac mole a moles total right sum  $n_1 \text{ all } x_{n_1}$  mole fraction of substance  
 a is often denoted by  $x_a$  and

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 mole fraction is not dependent on the temperature

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 present is equal to one  $x_a + x_b + \dots$  1 when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is  
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 total number of molecules of all the components in the solution

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 mass  $\text{C}_6\text{H}_6$  12.61 g mol 78 g mol 1 hence  $\text{C}_6\text{H}_6$  s number of moles  $\frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$  molar mass of carbon tet  $\text{CCl}_4$  1  
 2.4355 154 g mol 1

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 g of water calculate the molarity and molality of the solution given  $n_1 \text{ H}_1 \text{ C}_1 \text{ O}_1$  16 given mass of solute urea 11 11 g mass  
 of solvent water 100 g 0.1 kg to find molarity of solution solution

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web since the mole fraction of  $\text{O}_2$  is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.210 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature

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