

The curve in Figure 16.11 of the text shows how the vibrational heat capacity depends on temperature. Note that even when the temperature is only slightly above the vibrational temperature, the heat capacity is close to its equipartition value.

The total heat capacity of a molecular substance is the sum of each contribution (Figure 16.12 of the text). When equipartition is valid (when the temperature is well above the characteristic temperature of the mode  $T \gg \theta_M$ ) we can estimate the heat capacity by counting the numbers of modes that are active. In gases, all three translational modes are always active and contribute  $\frac{3}{2}R$  to the molar heat capacity. If we denote the number of active rotational modes by  $\nu_R^\ddagger$  (so for most molecules at normal temperatures  $\nu_R^\ddagger = 2$  for linear molecules, and 3 for non-linear molecules), then

$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

The equipartition value  $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R$  is reached when the temperature is high enough so that  $T \gg \theta_M$  for all the modes. However, if the temperature is lower, some of the modes are not fully excited and their contribution to the heat capacity is less than the equipartition value. The total heat capacity is then less than the equipartition value. The total heat capacity is then less than the equipartition value. The total heat capacity is then less than the equipartition value.

### Solutions to exercises

The total heat capacity is given by  $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$ . When equipartition is valid, the total heat capacity is  $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R$ .

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When equipartition is valid, the total heat capacity is  $C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R$ .

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The total heat capacity is

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$$C_{V,m} = \frac{3}{2}R + \frac{\nu_R^\ddagger}{2}R + \dots$$

When equipartition is valid,

# Physical Chemistry 9th Edition Instructors Solution Guide

**Katrin Zwirgmaier**



## **Physical Chemistry 9th Edition Instructors Solution Guide:**

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1. This book is structured into several chapters, namely:
  - Chapter 1: Introduction to Physical Chemistry 9th Edition Instructors Solution Guide
  - Chapter 2: Essential Elements of Physical Chemistry 9th Edition Instructors Solution Guide
  - Chapter 3: Physical Chemistry 9th Edition Instructors Solution Guide in Everyday Life
  - Chapter 4: Physical Chemistry 9th Edition Instructors Solution Guide in Specific Contexts
  - Chapter 5: Conclusion
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3. In chapter 2, the author will delve into the foundational concepts of Physical Chemistry 9th Edition Instructors Solution Guide. This chapter will elucidate the essential principles that need to be understood to grasp Physical Chemistry 9th Edition Instructors Solution Guide in its entirety.
4. In chapter 3, the author will examine the practical applications of Physical Chemistry 9th Edition Instructors Solution Guide in daily life. The third chapter will showcase real-world examples of how Physical Chemistry 9th Edition Instructors Solution Guide can be effectively utilized in everyday scenarios.
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This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Physical Chemistry 9th Edition Instructors Solution Guide.

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