

# Acids & Bases HOMEWORK

## Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- ~~HCl~~ **hydrochloric acid**
- ~~HCl~~ **hydrochloric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~HNO<sub>3</sub>~~ **nitric acid**
- ~~Sulfuric acid~~ **H<sub>2</sub>SO<sub>4</sub>**
- ~~potassium hydroxide~~ **KOH**
- ~~Sodium hydroxide~~ **NaOH**
- ~~calcium hydroxide~~ **Ca(OH)<sub>2</sub>**
- ~~ammonium hydroxide~~ **NH<sub>4</sub>OH**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.  
HClO<sub>4</sub> is STRONG with a lower cell completely dissociate in water

- HCl     
  HF     
  HClO<sub>2</sub>     
  Mg(OH)<sub>2</sub>     
  HNO<sub>3</sub>     
  H<sub>2</sub>SO<sub>4</sub>     
  NH<sub>4</sub>OH     
  NaOH

- What is the formula for hydronium? **H<sub>3</sub>O<sup>+</sup>**
- H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> will react with water and lose a H<sup>+</sup> ion in the process.
  - Write a balanced equation for this reaction. **H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> + H<sub>2</sub>O ⇌ H<sub>3</sub>O<sup>+</sup> + HC<sub>2</sub>O<sub>4</sub><sup>-</sup>**
  - Would you expect a large or small equilibrium constant for this reaction? **small, because H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is a weak acid.**
  - Would you define H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
  - In your written equation, is H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
  - Identify the conjugate base of H<sub>2</sub>C<sub>2</sub>O<sub>4</sub>. **HC<sub>2</sub>O<sub>4</sub><sup>-</sup>**
  - Identify H<sub>2</sub>C<sub>2</sub>O<sub>4</sub> as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**
- F<sup>-</sup> will react with water and gain a H<sup>+</sup> ion in the process.
  - Write a balanced equation for this reaction. **F<sup>-</sup> + H<sub>2</sub>O ⇌ HF + OH<sup>-</sup>**
  - Would you expect a large or small equilibrium constant for this reaction? **small, because F<sup>-</sup> is a weak base.**
  - Would you define F<sup>-</sup> as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H<sup>+</sup> ion.**
  - In your written equation, is F<sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
  - Identify the conjugate acid of F<sup>-</sup>. **HF**
  - Identify F<sup>-</sup> as polyprotic, amphoteric, or neither. **neither.**
- HCO<sub>3</sub><sup>-</sup> will react with water and lose a H<sup>+</sup> ion in the process.
  - Write a balanced equation for this reaction. **HCO<sub>3</sub><sup>-</sup> + H<sub>2</sub>O ⇌ H<sub>2</sub>O<sup>+</sup> + CO<sub>3</sub><sup>2-</sup>**
  - Would you expect a large or small equilibrium constant for this reaction? **small, because HCO<sub>3</sub><sup>-</sup> is a weak acid.**
  - Would you define HCO<sub>3</sub><sup>-</sup> as a proton donor or a proton acceptor? **proton donor, because it is losing a H<sup>+</sup> ion.**
  - In your written equation, is HCO<sub>3</sub><sup>-</sup> acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
  - Identify the conjugate base of HCO<sub>3</sub><sup>-</sup>. **CO<sub>3</sub><sup>2-</sup>**
  - Identify HCO<sub>3</sub><sup>-</sup> as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can act like an acid).**

# 19 3 Acids And Bases Answer Key Chemistry

**M Carnoy**



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