

## Chapter 9: Chemical Quantities

1. Solve stoichiometry problems using grams  $\rightarrow$  moles  $\rightarrow$  moles  $\rightarrow$  grams calculations. Examples below:

a. Calculate moles/grams of  $(\text{NH}_4)_2\text{SO}_4$  produced when 30.0 moles of  $\text{NH}_3$  react with  $\text{H}_2\text{SO}_4$ .



30.0 mol x moles

$$30.0 \text{ moles} \times 1 \text{ mol}/2 \text{ mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ grams/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b.  $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$ . How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams  $\rightarrow$  Moles  $\rightarrow$  Moles  $\rightarrow$  Grams

$$100. \text{ g Na}_2\text{CO}_3 \times 1 \text{ mol}/106.0 \text{ g} \times 2 \text{ mol NaNO}_3 / 1 \text{ mol Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g NaNO}_3 \times 1 \text{ mol}/85.00 \text{ g} \times 2 \text{ mol} / 2 \text{ mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of  $\text{NaNO}_3$  if 100.0 grams of  $\text{Na}_2\text{CO}_3$  are used in the reaction? What is the percent yield of  $\text{NaNO}_3$  if 130.5 grams of  $\text{NaNO}_3$  are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield}/\text{Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

## Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? *Absorb a photon of energy.*

How do they lose energy to return to their ground state? *Release a photon of same energy as that absorbed.*

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? *Specific wavelengths of light seen in emission spectra for atoms (page 285-86)*

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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**An Experimental Study of the Iowa Placement Examinations** Lawrence Williams Miller, University of Iowa, 1930  
*Circular of Information* USA. Bureau of Education, 1886      **Oxford University Gazette** University of Oxford, 1916  
**Program** Bryn Mawr College, 1905      British Medical Journal, 1913      **Contagious Diseases of domesticated Animals** D. E. Salmon, 1884      **Technical Education, an Economic Necessity** Victor Clifton Alderson, 1885      **The Universal tutor** London gen. corresp. coll, 1894      **The Lancet**, 1888      **The Prospective Teacher's Grasp of High School Content in His Probable Fields of Instruction** Russell E. Jonas, University of Iowa, 1928      Testing and Evaluation for the Sciences in the Secondary School William D. Hedges, 1966      *Bryn Mawr College Calendar* Bryn Mawr College, 1924      *A Project in Student Personnel Service Designed to Facilitate Each Student's Achievement at the Level of His Ability* Lonzo Jones, 1930      Calendar Bryn Mawr College, 1885      *School Science and Mathematics*, 1921      **Science Tests and Reviews** Buros Center, 1975 Science Tests and Reviews consisting of science sections of the first seven MMYs and Tests in Print II includes 217 original test reviews written by 81 specialists 18 excerpted test reviews 270 references on the construction use and validity of specific tests a bibliography on in print science tests references for specific tests cumulative name indexes for specific tests with references a publishers directory title index name index and a scanning index The 97 tests covered fall into the following categories 23 general 14 biology 35 chemistry 3 geology 6 miscellaneous and 16 physics

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