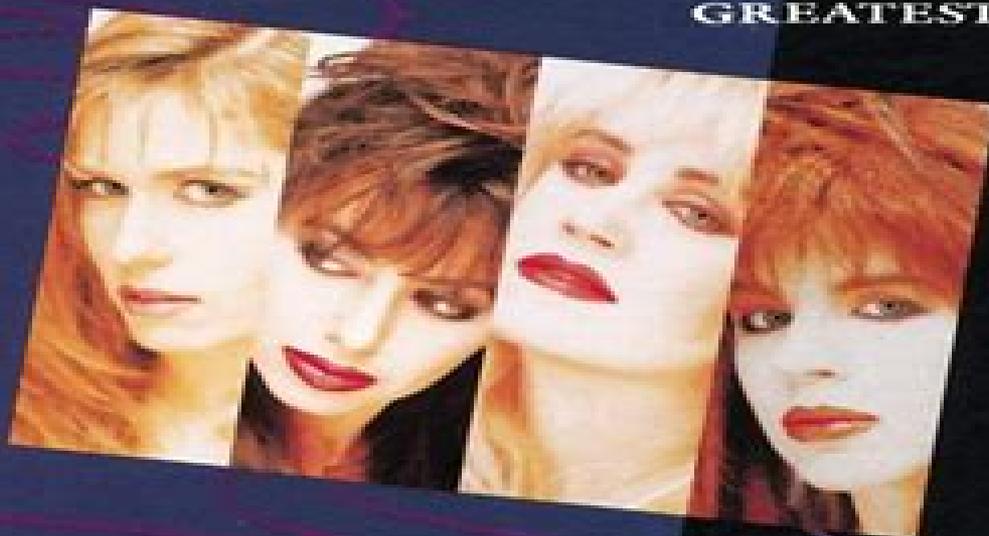


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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and Δt is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of *guided inquiry designing the ideal hand warmer* - Dec 27 2022

web the heat of reaction ΔH_{soln} is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as $NaOH(s) \rightarrow NaOH(aq)$ $\Delta H_{soln} = -44.2 \text{ kJ mol}^{-1}$ rxn write equations to similarly represent the dissolving process for

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