

## Chemistry 2<sup>nd</sup> Semester Final Review **KEY**

\*This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.\*

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

*Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)*

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)  
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$   
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$   
Since all units are standard (L, atm), we do not need to convert the units before proceeding.  
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
T goes away (not mentioned in problem)  
 $P_1V_1=P_2V_2$   
Isolate for P<sub>2</sub>:  $P_2=(P_1V_1)/V_2$   
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$  (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)  
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$   
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$   
We need to convert from Celsius to Kelvins before proceeding.  
 $T_1=34.75 + 273=307.75 \text{ K}$   
 $T_2=78.41 + 273=351.41 \text{ K}$   
 $(P_1V_1)/T_1=(P_2V_2)/T_2$   
Get rid of P  
 $V_1/T_1=V_2/T_2$   
Isolate V<sub>2</sub>:  $V_2=(V_1/T_1)T_2$   
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?  
The combined gas law won't work here. You only have one condition and nothing changes.  
 $PV=nRT$  will work. (n= moles and R= 0.08206 L atm/K mol)  
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$  OOPS we need to convert from grams to moles!,  $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$   
n: Hydrogen gas is diatomic!  
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$   
Isolate V:  $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

# 2nd Semester Chemistry Final Exam Study Guide

**Marcel A. Müller**



## **2nd Semester Chemistry Final Exam Study Guide:**

## **2nd Semester Chemistry Final Exam Study Guide** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**2nd Semester Chemistry Final Exam Study Guide**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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