

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H₂SO₄ **sulfuric acid**
- HNO₃ **nitric acid**
- HClO₄ **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **CaOH**
- calcium hydroxide **Ca(OH)₂**
- ammonium hydroxide **Li(OH)₄**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
(HClO₄ & HNO₃) will dissociate with a lower pH completely dissociate in water

- HCl
 HF
 H₂S
 Mg(OH)₂
 H₂SO₄
 H₂CO₃
 NH₄OH
 HClO₄

3. What is the formula for hydrosulfuric acid? **H₂S**

4. H₂CO₃ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction: **H₂CO₃ + H₂O ⇌ H₂O⁺ + HCO₃⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂CO₃ is a weak acid.**
- Would you define H₂CO₃ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂CO₃ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of H₂CO₃: **HCO₃⁻**
- Identify H₂CO₃ as polyprotic, amphoteric, or neither: **polyprotic, because it contains multiple H's to give.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for the reaction: **F⁻ + H₂O ⇌ HF + OH⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
- Identify the conjugate acid of F⁻: **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither: **neither.**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction: **HSO₄⁻ + H₂O ⇌ H₂O⁺ + SO₄²⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of HSO₄⁻: **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither: **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can give out H's too).**

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Léo Pomar



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Table of Contents 19 3 Acids And Bases Answer Key Chemistry

1. Understanding the eBook 19 3 Acids And Bases Answer Key Chemistry
 - The Rise of Digital Reading 19 3 Acids And Bases Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 3 Acids And Bases Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 3 Acids And Bases Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 3 Acids And Bases Answer Key Chemistry
 - Personalized Recommendations
 - 19 3 Acids And Bases Answer Key Chemistry User Reviews and Ratings
 - 19 3 Acids And Bases Answer Key Chemistry and Bestseller Lists
5. Accessing 19 3 Acids And Bases Answer Key Chemistry Free and Paid eBooks
 - 19 3 Acids And Bases Answer Key Chemistry Public Domain eBooks
 - 19 3 Acids And Bases Answer Key Chemistry eBook Subscription Services

- 19 3 Acids And Bases Answer Key Chemistry Budget-Friendly Options
- 6. Navigating 19 3 Acids And Bases Answer Key Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - 19 3 Acids And Bases Answer Key Chemistry Compatibility with Devices
 - 19 3 Acids And Bases Answer Key Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 3 Acids And Bases Answer Key Chemistry
 - Highlighting and Note-Taking 19 3 Acids And Bases Answer Key Chemistry
 - Interactive Elements 19 3 Acids And Bases Answer Key Chemistry
- 8. Staying Engaged with 19 3 Acids And Bases Answer Key Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 3 Acids And Bases Answer Key Chemistry
- 9. Balancing eBooks and Physical Books 19 3 Acids And Bases Answer Key Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 3 Acids And Bases Answer Key Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 19 3 Acids And Bases Answer Key Chemistry
 - Setting Reading Goals 19 3 Acids And Bases Answer Key Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 19 3 Acids And Bases Answer Key Chemistry
 - Fact-Checking eBook Content of 19 3 Acids And Bases Answer Key Chemistry
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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