

## Chemistry:

### CHAPTER 1

- Charge of electron  $\approx -1.6 \times 10^{-19}$ ; Avogadro's #  $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$ ;  $E = qV$  (energy = charge on e- x electron volts);  $c = \lambda\nu$  ( $c = 3 \times 10^8$ )
- **Heisenberg uncertainty** – impossible to determine momentum and position simultaneously
- **Pauli exclusion** – no electrons have same quantum #'s
- Max e- in a shell  $\approx 2n^2$ , max e- in subshell  $\approx 4l+2$ 
  - Possible values for l  $\approx 0$  to  $n-1$ , possible values for ml  $\approx -l$  to  $+l$
- s can hold 2 e-, p  $\approx 6$ , d  $\approx 10$ , f  $\approx 14$
- **Aufbau principle** (building up) – lowest  $n+l$  will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- **Hund's** – half filled orbitals with parallel spins (bus seats)
- **Paramagnetic** – unpaired parallel spin e-, attracted to magnetic field
  - **Diamagnetic** – paired e-, repelled by magnetic field
  - **Ferromagnetic** – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

### CHAPTER 2

- **Z eff** increases from left to right, constant up and down
- **Ionic radius** is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form **hydration complexes**
- **Atomic radius** has the opposite trend of all others



### CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- **Coordinate covalent bond** – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic  $\approx$  EN difference greater than 1.7, polar covalent  $\approx$  0.5 to 1.7, nonpolar  $\approx$  less than 0.5
  - Ionic are stronger than covalent (which are both stronger than IM forces)
    - Van der Waals occurs more with close proximity and large molecules
- **Dipole moment**  $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
  - **Formal charge = valence – nonbonding – ½ bonding**
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

# Section 2 nuclear Chemistry Study Guide Key

**LL Leslie**



**Section 2 nuclear Chemistry Study Guide Key:**

This book delves into Section 2nuclear Chemistry Study Guide Key. Section 2nuclear Chemistry Study Guide Key is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Section 2nuclear Chemistry Study Guide Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
    - Chapter 1: Introduction to Section 2nuclear Chemistry Study Guide Key
    - Chapter 2: Essential Elements of Section 2nuclear Chemistry Study Guide Key
    - Chapter 3: Section 2nuclear Chemistry Study Guide Key in Everyday Life
    - Chapter 4: Section 2nuclear Chemistry Study Guide Key in Specific Contexts
    - Chapter 5: Conclusion
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  3. In chapter 2, this book will delve into the foundational concepts of Section 2nuclear Chemistry Study Guide Key. The second chapter will elucidate the essential principles that need to be understood to grasp Section 2nuclear Chemistry Study Guide Key in its entirety.
  4. In chapter 3, the author will examine the practical applications of Section 2nuclear Chemistry Study Guide Key in daily life. The third chapter will showcase real-world examples of how Section 2nuclear Chemistry Study Guide Key can be effectively utilized in everyday scenarios.
  5. In chapter 4, this book will scrutinize the relevance of Section 2nuclear Chemistry Study Guide Key in specific contexts. This chapter will explore how Section 2nuclear Chemistry Study Guide Key is applied in specialized fields, such as education, business, and technology.
  6. In chapter 5, the author will draw a conclusion about Section 2nuclear Chemistry Study Guide Key. The final chapter will summarize the key points that have been discussed throughout the book.
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## Table of Contents Section 2nuclear Chemistry Study Guide Key

1. Understanding the eBook Section 2nuclear Chemistry Study Guide Key
  - The Rise of Digital Reading Section 2nuclear Chemistry Study Guide Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 2nuclear Chemistry Study Guide Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 2nuclear Chemistry Study Guide Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2nuclear Chemistry Study Guide Key
  - Personalized Recommendations
  - Section 2nuclear Chemistry Study Guide Key User Reviews and Ratings
  - Section 2nuclear Chemistry Study Guide Key and Bestseller Lists
5. Accessing Section 2nuclear Chemistry Study Guide Key Free and Paid eBooks
  - Section 2nuclear Chemistry Study Guide Key Public Domain eBooks
  - Section 2nuclear Chemistry Study Guide Key eBook Subscription Services
  - Section 2nuclear Chemistry Study Guide Key Budget-Friendly Options
6. Navigating Section 2nuclear Chemistry Study Guide Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 2nuclear Chemistry Study Guide Key Compatibility with Devices
  - Section 2nuclear Chemistry Study Guide Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 2nuclear Chemistry Study Guide Key
  - Highlighting and Note-Taking Section 2nuclear Chemistry Study Guide Key
  - Interactive Elements Section 2nuclear Chemistry Study Guide Key

8. Staying Engaged with Section 2nuclear Chemistry Study Guide Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 2nuclear Chemistry Study Guide Key
9. Balancing eBooks and Physical Books Section 2nuclear Chemistry Study Guide Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 2nuclear Chemistry Study Guide Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Section 2nuclear Chemistry Study Guide Key
  - Setting Reading Goals Section 2nuclear Chemistry Study Guide Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 2nuclear Chemistry Study Guide Key
  - Fact-Checking eBook Content of Section 2nuclear Chemistry Study Guide Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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