

## Chemistry:

### CHAPTER 1

- Charge of electron  $\approx -1.6 \times 10^{-19}$ ; Avogadro's #  $\approx 6 \times 10^{23}$
- Mass number (A) on top of element, atomic number (Z) below
- $E = \frac{hc}{\lambda} = -R \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$ ;  $E = qV$  (energy = charge on e- x electron volts);  $c = \lambda \nu$  ( $c = 3 \times 10^8$ )
- Heisenberg uncertainty – impossible to determine momentum and position simultaneously
- Pauli exclusion – no electrons have same quantum #'s
- Max e- in a shell  $\approx 2n^2$ , max e- in subshell  $\approx 4l+2$ 
  - Possible values for l  $\approx 0$  to  $n-1$ , possible values for ml  $\approx -l$  to  $+l$
- s can hold 2 e-, p  $\approx 6$ , d  $\approx 10$ , f  $\approx 14$
- Aufbau principle (building up) – lowest  $n+l$  will fill first
- For anions add extra e-, for cations remove e- from subshells with highest n first
- Hund's – half filled orbitals with parallel spins (bus seats)
- Paramagnetic – unpaired parallel spin e-, attracted to magnetic field
  - Diamagnetic – paired e-, repelled by magnetic field
  - Ferromagnetic – unpaired e-, randomly oriented dipoles, become magnetized in field
- A faraday is equal to one mole of electric charge

### CHAPTER 2

- $Z_{\text{eff}}$  increases from left to right, constant up and down
- Ionic radius is smallest near metalloid line and increases out
- F is smallest, most EN, highest IE and EA; Cs is the largest and least EN
- Complex ions from transition metals associate with water to form hydration complexes
- Atomic radius has the opposite trend of all others



### CHAPTER 3

- C, N, O, F, Na, Mg always abide by octet rule; H, He, Li, Be, B can have incomplete; period 3 or higher can have expanded
- Coordinate covalent bond – both e- are contributed by one atom (lone pair attack, acid-base)
- Ionic  $\approx$  EN difference greater than 1.7, polar covalent  $\approx$  0.5 to 1.7, nonpolar  $\approx$  less than 0.5
  - Ionic are stronger than covalent (which are both stronger than IM forces)
    - Van der Waals occurs more with close proximity and large molecules
- Dipole moment  $\approx p = qd$
- Triple bonds are shorter and have more energy
- Major resonance contributor is more stable (less formal charge, negative on more EN atom)
  - Formal charge = valence – nonbonding –  $\frac{1}{2}$  bonding
- Electron geometry is when bonding and lone are equal, molecular geo is when lone repel more

# Section 2 nuclear Chemistry Study Guide Key

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