

Chemistry 2nd Semester Final Review **KEY**

This is meant as a guide. It does NOT contain every single thing you need to know for the final exam. Review notes, benchmark review sheets, homeworks, chapter study guides, etc.

Bring your book to school by Monday for turn in (seniors should bring book to bookstore by themselves ASAP)

Gas Laws (Ch. 13):

Subtopics: atmospheric pressure, Dalton's law of partial pressures, combined gas law, Ideal gas law, gas stoichiometry (22.4 L/mol for a gas at STP)

Practice questions:

- If a gas occupies 3.8 L at a pressure of 2.71 atm, what would the pressure be if the volume changes to 1.47 L? (Temp. and amt. of gas constant)
 $V_1=3.8 \text{ L}; P_1=2.71 \text{ atm}$
 $V_2=1.47 \text{ L}; P_2=? \text{ Atm}$
Since all units are standard (L, atm), we do not need to convert the units before proceeding.
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
T goes away (not mentioned in problem)
 $P_1V_1=P_2V_2$
Isolate for P₂: $P_2=(P_1V_1)/V_2$
 $P_2=(2.71 \text{ atm} \times 3.8 \text{ L})/(1.47 \text{ L})=7.0 \text{ atm}$ (L cancelled and it's 2 sig figs)
- A gas has a volume of 1.49 L at a temperature of 34.75 degrees Celsius. What would the volume be at 78.41 degrees Celsius? (pressure & amt. of gas constant)
 $V_1=1.49 \text{ L}, T_1=34.75 \text{ degrees Celsius}$
 $V_2=? \text{ L}, T_2=78.41 \text{ degrees Celsius}$
We need to convert from Celsius to Kelvins before proceeding.
 $T_1=34.75 + 273=307.75 \text{ K}$
 $T_2=78.41 + 273=351.41 \text{ K}$
 $(P_1V_1)/T_1=(P_2V_2)/T_2$
Get rid of P
 $V_1/T_1=V_2/T_2$
Isolate V₂: $V_2=(V_1/T_1)T_2$
 $V_2=(1.49 \text{ L}/307.75 \text{ K})(351.41 \text{ K})=1.70 \text{ L}$
- What volume is occupied by 8.47 g of hydrogen gas at 84.7 degrees Celsius and 1.04 atm?
The combined gas law won't work here. You only have one condition and nothing changes.
 $PV=nRT$ will work. (n = moles and R = 0.08206 L atm/K mol)
 $P=1.04 \text{ atm}, V=? \text{ L}, n=?$ OOPS we need to convert from grams to moles!, $T=84.7 \text{ degrees Celsius} + 273=357.7 \text{ K}$
 n : Hydrogen gas is diatomic!
 $8.47 \text{ g H}_2 \times (1 \text{ mol H}_2/2.02 \text{ g H}_2)=4.193069307 \text{ mol H}_2$
Isolate V: $V=(nRT)/P=(4.193069307 \text{ mol H}_2 \times 0.08206 \text{ L atm/K mol} \times 357.7 \text{ K})/1.04 \text{ atm}=118 \text{ L}$
- What volume is occupied by 56.75 g of oxygen gas at STP?

2nd Semester Chemistry Final Exam Review Answers

RS Peters



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Table of Contents 2nd Semester Chemistry Final Exam Review Answers

1. Understanding the eBook 2nd Semester Chemistry Final Exam Review Answers
 - The Rise of Digital Reading 2nd Semester Chemistry Final Exam Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Chemistry Final Exam Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Chemistry Final Exam Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Chemistry Final Exam Review Answers

- Personalized Recommendations
 - 2nd Semester Chemistry Final Exam Review Answers User Reviews and Ratings
 - 2nd Semester Chemistry Final Exam Review Answers and Bestseller Lists
5. Accessing 2nd Semester Chemistry Final Exam Review Answers Free and Paid eBooks
 - 2nd Semester Chemistry Final Exam Review Answers Public Domain eBooks
 - 2nd Semester Chemistry Final Exam Review Answers eBook Subscription Services
 - 2nd Semester Chemistry Final Exam Review Answers Budget-Friendly Options
 6. Navigating 2nd Semester Chemistry Final Exam Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 2nd Semester Chemistry Final Exam Review Answers Compatibility with Devices
 - 2nd Semester Chemistry Final Exam Review Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2nd Semester Chemistry Final Exam Review Answers
 - Highlighting and Note-Taking 2nd Semester Chemistry Final Exam Review Answers
 - Interactive Elements 2nd Semester Chemistry Final Exam Review Answers
 8. Staying Engaged with 2nd Semester Chemistry Final Exam Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2nd Semester Chemistry Final Exam Review Answers
 9. Balancing eBooks and Physical Books 2nd Semester Chemistry Final Exam Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2nd Semester Chemistry Final Exam Review Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 2nd Semester Chemistry Final Exam Review Answers
 - Setting Reading Goals 2nd Semester Chemistry Final Exam Review Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 2nd Semester Chemistry Final Exam Review Answers

- Fact-Checking eBook Content of 2nd Semester Chemistry Final Exam Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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