

BAGONG SILANG NATIONAL HIGH SCHOOL
1ST QUARTER EXAMINATION
PHYSICAL SCIENCE

NAME: _____
GRADE: _____

DATE: _____
SCORE: _____

Directions: Write the letter of the correct answer on the space provided.

- _____ 1. It is also known as nuclear fusion and the formation of new nuclei actions in the early stages of development of the universe.
a. Nucleosynthesis b. S-Process c. R-Process d. Singularity
- _____ 2-4. There are two categories of star, the average star and massive star. The life cycle of these two stars are common from 1st to 3rd stages namely nebula, protostar, and main sequence star. What will be the next stages in the life of an average star until it reaches its death?
a. Red super giant → Supernova → Black hole → neutron star 0
b. Red giant → white dwarf → black dwarf 3
c. Red giant → white dwarf → black hole 2
d. Red supergiant → white dwarf → black dwarf 1
- _____ 5. He synthesized element with atomic number 43 using a linear particle accelerator.
a. Ernest Rutherford b. Ernest Lawrence c. Dmitri Mendeleev d. John Dalton
- _____ 6. The following are light elements EXCEPT;
a. helium b. hydrogen c. lithium d. iron
- _____ 7. Which of the following molecule is nonpolar?
a. NaCl b. HCl c. CO₂ d. NH₃
- _____ 8. What type of chemical bond holds the atoms of water molecule together?
a. Hydrogen bond b. Ionic bond c. Polar covalent bond d. Nonpolar covalent bond
- _____ 9. Which of the following IMFAs is considered as the weakest?
a. H-bonding b. Ion-dipole c. Dipole-dipole d. London forces
- _____ 10. Which of the following statements is TRUE about nonpolar molecules?
a. high boiling point b. high melting point c. low surface tension d. low vapor pressure
- _____ 11. Forces that are present between an ion and a polar molecule
a. ion-dipole b. dipole-dipole c. dispersion forces d. hydrogen bonding
- _____ 12. H-bonding forms when the substances involved are polar and there are _____.
a. H-atoms of a molecule interact with to O, N, F of other molecules.
b. C-atoms attached to O, N, F.
c. central atoms with O, N, F as attached atoms.
d. unshared pair of electrons in the central atom.
- _____ 13. The amount of energy required to stretch or increase the surface of a liquid by a unit area
a. specific heat b. surface tension c. vapor pressure d. heat of vaporization
- _____ 14. Liquids can form spherical elastic film to minimize surface area. What intermolecular force is responsible for the formation of this film in water?
a. H-bonding b. ion-induced dipole c. dipole-induced dipole d. London dispersion force
- _____ 15. Which group of biomolecules carries and passes on the hereditary information of the organisms?
a. carbohydrates b. lipids c. nucleic acids d. proteins
- _____ 16. What are the elements that make up carbohydrates?
a. C and H b. C, H and N c. C, H and O d. C, H, O and N
- _____ 17-19. Is carbohydrates one of the most important biological macromolecules needed by a human body?
a. Yes, because it is the primary source of energy. 3
b. Yes, because it can be a storage of energy for a long period of time. 1
c. Yes, because it can be found in rice, cereals, fruits and pastas. 2
d. Yes, because it carries genetic information. 0
- _____ 20. The measure of how fast is the change in the concentration of the reactants or products in a chemical reaction.
a. Activation Energy b. Activation Rate c. Chemical Reaction d. Concentration
- _____ 21. A chemical change that occurs when two or more substances combine to form a new substance.
a. Activation Energy b. Chemical Kinetics c. Chemical Reaction d. Product

Section 2review Answer Key Physical Science

Robert R. Redfield



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1. Understanding the eBook Section 2review Answer Key Physical Science
 - The Rise of Digital Reading Section 2review Answer Key Physical Science
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 2review Answer Key Physical Science
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 2review Answer Key Physical Science
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 2review Answer Key Physical Science
 - Personalized Recommendations
 - Section 2review Answer Key Physical Science User Reviews and Ratings
 - Section 2review Answer Key Physical Science and Bestseller Lists
5. Accessing Section 2review Answer Key Physical Science Free and Paid eBooks

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 - Section 2review Answer Key Physical Science eBook Subscription Services
 - Section 2review Answer Key Physical Science Budget-Friendly Options
6. Navigating Section 2review Answer Key Physical Science eBook Formats
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 - Section 2review Answer Key Physical Science Compatibility with Devices
 - Section 2review Answer Key Physical Science Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Section 2review Answer Key Physical Science
 - Highlighting and Note-Taking Section 2review Answer Key Physical Science
 - Interactive Elements Section 2review Answer Key Physical Science
8. Staying Engaged with Section 2review Answer Key Physical Science
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 2review Answer Key Physical Science
9. Balancing eBooks and Physical Books Section 2review Answer Key Physical Science
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 2review Answer Key Physical Science
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 2review Answer Key Physical Science
- Setting Reading Goals Section 2review Answer Key Physical Science
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 2review Answer Key Physical Science
- Fact-Checking eBook Content of Section 2review Answer Key Physical Science
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web the rate law for a particular reaction is rate $k x^2$ in an experiment the initial rate of the reaction is determined to be 0.080 mol l s when the initial concentration of x is

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web a rate law shows how the rate of a chemical reaction depends on reactant concentration for a reaction such as $a \rightarrow \text{products}$ the rate law generally has the form rate $k a^n$

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web the reaction orders in a rate law describe the mathematical dependence of the rate on reactant concentrations referring to the generic rate law above the reaction is m

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web each reactant in the rate limiting step of a reaction is assigned an order typically zeroth 1st or 2nd the order is the sum of all individual orders the rate of a reaction is

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web the rate law or differential rate law relates the rate of a reaction to the concentration or pressure of the reactants the rate of a reaction is proportional to the concentration or

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web the absorbance of a particular wavelength of light by a solution 2 in this experiment you will conduct the reaction between solutions of potassium iodide and iron iii chloride

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web 1 use table 1 to determine the reaction order for hcl 2 use table 2 to determine the reaction order for na₂s₂o₃ remember

you want to see what happens to the reaction

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