

9 10 11 12

DOUBLE UNI-KNOT

An easy to tie and reliable knot to join two lines.

1. Mono to braid: 5x in mono, 5x in braid. 4 Wraps

2. 4 Wraps 1st Uni-knot

4 Wraps 2nd Uni-knot

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Leader

Both lines through 2x

UNI KNOT

1. 2. 3. 4. 5. 6. 7. 8. 9. 10. 11. 12.

1. 2. 3. 4. 5. 6. 7. 8. 9. 10. 11. 12.

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1. 5 - 6 wraps

2. 3.

1. 2. 3. 4.

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Fishermans Ultimate Knot Guide

Kathleen Armour



Fishermans Ultimate Knot Guide:

Fishermans Ultimate Knot Guide Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has are more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Fishermans Ultimate Knot Guide**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve into the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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 $2 \text{ CO}_2 \text{ (g)} + \text{H}_2\text{O (g)} \rightarrow \text{C}_2\text{H}_2 \text{ (g)} + \frac{5}{2} \text{O}_2 \text{ (g)}$ $\Delta h = 226 \text{ kJ}$
 $\text{C}_2\text{H}_2 \text{ (g)} + 2 \text{ H}_2 \text{ (g)} \rightarrow \text{C}_2\text{H}_6 \text{ (g)}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O}_2 \text{ (g)}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ (g)} + \frac{7}{2} \text{O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (g)}$ $\Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4 \text{ (g)} + 3 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 2 \text{ H}_2\text{O (l)}$ $\Delta h = 1411 \text{ kJ}$ mole $\text{C}_2\text{H}_6 \text{ (g)} + 7 \text{ O}_2 \text{ (g)} \rightarrow 2 \text{ CO}_2 \text{ (g)} + 3 \text{ H}_2\text{O (l)}$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta h = 285 \text{ kJ}$ mole

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = 55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $2\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ $\text{H}_2\text{O}(\text{l}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Cl}^-(\text{aq})$

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