

**EXPERIMENT NO: 1**

**DATE:**    /    /    .

**AIM - Determination of the pH of the given water Sample.**

**APPARATUS –**

- i. PH meter along with electrodes
- ii. Buffer solution
- iii. Thermometer
- iv. PH paper

**PRINCIPLE –**

pH of water is a measure of amount of hydrogen ions that is present in the water. It determines if the water is alkaline or acidic in nature. pH stands for potential of hydrogen, mathematically as

$$\text{pH} = - \log [\text{H}^+]$$

Potentiometric pH meters measure the voltage between two electrodes varying linearly with the pH of the solution in which it is immersed and display the result converted into the corresponding pH value. It comprises of a simple electronic amplifier and a pair of electrodes and some form of display calibrated in pH units. It usually has a **glass electrode** and a **reference electrode** contained within a single probe or a combination electrode, a bulb containing the sensor at the bottom specifically designed to be selective to hydrogen-ion concentration. The electrodes, or probes, are inserted into the solution to be tested.

In order to accurately measure the potential difference between the two sides of the glass membrane reference electrode, which has a stable and well-known electrode potential typically a **silver chloride electrode** or **calomel electrode** are required on each side of the membrane. Their purpose is to measure changes in the potential on their respective side.

Calibration is performed with at least two standard **buffer solutions** that span the range of pH values to be measured.

The magnitude of the electrochemical potential across the glass bulb is linearly related to the pH according to the Nernst equation.

# Nd Sem Chemistry Lab Manual For Rtu

**Albert A Gayle**



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