

Lectures 21-22

Solid state materials. Electronic structure and conductivity

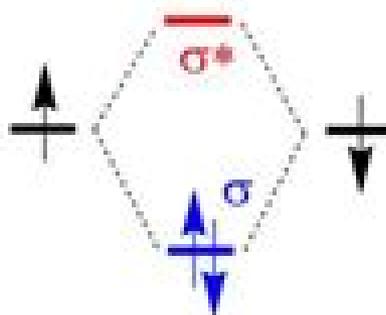
1) Band theory

- The electronic structure of **solids** can also be described by MO theory.
- A solid can be considered as a **supermolecule**.
- One mole of atoms (N_A), each with X orbitals in the valence shell contributes X moles of atomic orbitals producing X moles of MO's.

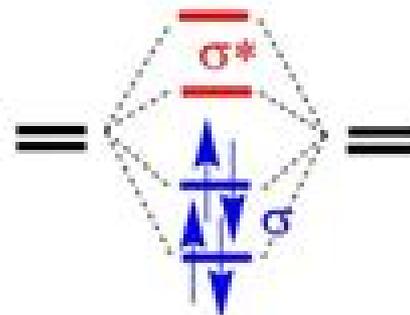
Consider qualitatively bonding between N metal atoms of ns^1 configuration (Li, Na etc) arranged in a *chain*; $N = 2, 4, N_A$. Assume that $X=1$ for simplicity.

- In the case of $N \sim N_A$ atoms they form not **bonds** but **bands**.
- The band appearing in the *bonding* region is called **valence band**. The *antibonding* region is called **conduction band**.
- In the case of metals the **valence** and **conduction** bands are immediately adjacent.

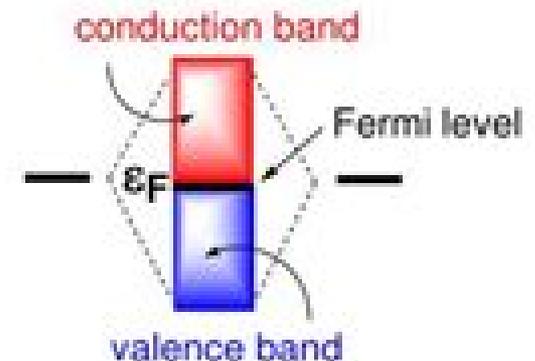
$N = 2$ Li atoms



4 Li atoms



N_A Li atoms



Atomic And Electronic Structure Of Solids

K Payea



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